

Penny Lab at home

Name: _____ partner: _____

The United States penny has been made of 95% copper from 1980 - 1982. However, the price of copper had risen so much that the U.S. Treasury decided to change the penny's intended alloy. Hence, it was intended that the composition of the penny, for all 1983 pennies have been made of mostly zinc, with a thin copper layer on the outside. In this lab, you will determine the % zinc and % copper in a penny. To do this you will take advantage of the activity series of metals, and perform a single replacement reaction on the penny.

Procedure

1. Obtain a clean, unscratched penny (post 1982, the newer the better). You and your lab partner will each need different years so you can tell them apart! Record the years in the data table.
2. Using a file, carefully scratch the edge of your penny in two places, until the zinc core is showing.
3. Weigh each penny. Record the mass.
4. Place each penny in a beaker with ca. 50 mL of HCl. Label the beaker with your names. Remove the bubbling around the filed area. Set beaker aside in designated area.

— Day Two —

5. Decant HCl and rinse penny w/ water. Carefully remove penny & dry w/ paper towel.
6. Once dry, weigh copper shell and record mass.

Calculations / Questions

1. How much zinc was lost in penny 1? _____ g (mass - mass)

2. How much zinc was lost in penny 2? _____ g (mass - mass)

3. What is the % Zn in penny 1?

4. What is the % Zn in penny 2?

5. What is the average % Zn?

6. The actual % Zn in pennies is 2.43%. Calculate % error. $(\text{actual} - \text{experimental}) \div \text{actual} \times 100$

7. Write the complete, balanced equation for the chemical reaction you did.

8. Assuming the average penny contains 2.50 g zinc, how much HCl is required to react the all of this zinc, without leaving any EXCESS HCl? (hint: look at stoichiometry section, w/ your teacher's advice before attempting this)



	Penny 1	Penny 2
Your Name:		
Penny Year:		
Initial Mass (g):		
Final Mass (g):		

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