

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The sample is a solution of a chloride salt. The sample is weighed and then dissolved in water. The solution is then treated with a solution of silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a fume hood to avoid exposure to the silver chloride precipitate, which is a strong irritant. The experiment is performed in a fume hood to avoid exposure to the silver nitrate solution, which is a strong oxidizing agent and can cause skin irritation. The experiment is performed in a fume hood to avoid exposure to the silver chloride precipitate, which is a strong irritant. The experiment is performed in a fume hood to avoid exposure to the silver nitrate solution, which is a strong oxidizing agent and can cause skin irritation. The experiment is performed in a fume hood to avoid exposure to the silver chloride precipitate, which is a strong irritant.

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Procedure

The procedure for this experiment is as follows: 1. Weigh a sample of the chloride salt. 2. Dissolve the sample in water. 3. Add a solution of silver nitrate to the solution. 4. Filter the precipitate. 5. Dry the precipitate. 6. Weigh the precipitate. 7. Calculate the mass of chloride in the original sample.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.100 g	0.100 g
Mass of precipitate	0.150 g	0.150 g
Mass of chloride in sample	0.050 g	0.050 g

The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.150 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.037 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.150 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.037 \text{ g}$

Analysis Of Chloride Lab Report

G Psacharopoulos



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