

Balancing Equations Worksheet - Solutions

- 1) $\underline{1} \text{ H}_3\text{PO}_4 + \underline{3} \text{ KOH} \rightarrow \underline{1} \text{ K}_3\text{PO}_4 + \underline{3} \text{ H}_2\text{O}$
- 2) $\underline{6} \text{ K} + \underline{1} \text{ B}_2\text{O}_3 \rightarrow \underline{3} \text{ K}_2\text{O} + \underline{2} \text{ B}$
- 3) $\underline{1} \text{ HCl} + \underline{1} \text{ NaOH} \rightarrow \underline{1} \text{ NaCl} + \underline{1} \text{ H}_2\text{O}$
- 4) $\underline{10} \text{ Na} + \underline{2} \text{ NaNO}_3 \rightarrow \underline{6} \text{ Na}_2\text{O} + \underline{1} \text{ N}_2$
- 5) $\underline{4} \text{ C} + \underline{1} \text{ S}_8 \rightarrow \underline{4} \text{ CS}_2$
- 6) $\underline{4} \text{ Na} + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ Na}_2\text{O}$
- 7) $\underline{2} \text{ N}_2 + \underline{5} \text{ O}_2 \rightarrow \underline{2} \text{ N}_2\text{O}_5$
- 8) $\underline{2} \text{ H}_3\text{PO}_4 + \underline{3} \text{ Mg(OH)}_2 \rightarrow \underline{1} \text{ Mg}_3(\text{PO}_4)_2 + \underline{6} \text{ H}_2\text{O}$
- 9) $\underline{2} \text{ NaOH} + \underline{1} \text{ H}_2\text{CO}_3 \rightarrow \underline{1} \text{ Na}_2\text{CO}_3 + \underline{2} \text{ H}_2\text{O}$
- 10) $\underline{1} \text{ KOH} + \underline{1} \text{ HBr} \rightarrow \underline{1} \text{ KBr} + \underline{1} \text{ H}_2\text{O}$
- 11) $\underline{4} \text{ Na} + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ Na}_2\text{O}$
- 12) $\underline{2} \text{ Al(OH)}_3 + \underline{3} \text{ H}_2\text{CO}_3 \rightarrow \underline{1} \text{ Al}_2(\text{CO}_3)_3 + \underline{6} \text{ H}_2\text{O}$
- 13) $\underline{16} \text{ Al} + \underline{3} \text{ S}_8 \rightarrow \underline{8} \text{ Al}_2\text{S}_3$
- 14) $\underline{6} \text{ Cs} + \underline{1} \text{ N}_2 \rightarrow \underline{2} \text{ Cs}_3\text{N}$
- 15) $\underline{1} \text{ Mg} + \underline{1} \text{ Cl}_2 \rightarrow \underline{1} \text{ MgCl}_2$
- 16) $\underline{10} \text{ Rb} + \underline{2} \text{ RbNO}_3 \rightarrow \underline{6} \text{ Rb}_2\text{O} + \underline{1} \text{ N}_2$
- 17) $\underline{2} \text{ C}_6\text{H}_6 + \underline{15} \text{ O}_2 \rightarrow \underline{12} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O}$
- 18) $\underline{1} \text{ N}_2 + \underline{3} \text{ H}_2 \rightarrow \underline{2} \text{ NH}_3$
- 19) $\underline{2} \text{ C}_{10}\text{H}_{22} + \underline{31} \text{ O}_2 \rightarrow \underline{20} \text{ CO}_2 + \underline{22} \text{ H}_2\text{O}$
- 20) $\underline{1} \text{ Al(OH)}_3 + \underline{3} \text{ HBr} \rightarrow \underline{1} \text{ AlBr}_3 + \underline{3} \text{ H}_2\text{O}$
- 21) $\underline{2} \text{ CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \underline{13} \text{ O}_2 \rightarrow \underline{8} \text{ CO}_2 + \underline{10} \text{ H}_2\text{O}$
- 22) $\underline{1} \text{ C}_3\text{H}_8 + \underline{5} \text{ O}_2 \rightarrow \underline{3} \text{ CO}_2 + \underline{4} \text{ H}_2\text{O}$
- 23) $\underline{3} \text{ Li} + \underline{1} \text{ AlCl}_3 \rightarrow \underline{3} \text{ LiCl} + \underline{1} \text{ Al}$
- 24) $\underline{2} \text{ C}_3\text{H}_8 + \underline{7} \text{ O}_2 \rightarrow \underline{4} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O}$
- 25) $\underline{3} \text{ NH}_4\text{OH} + \underline{1} \text{ H}_3\text{PO}_4 \rightarrow \underline{1} \text{ (NH}_4)_3\text{PO}_4 + \underline{3} \text{ H}_2\text{O}$
- 26) $\underline{3} \text{ Rb} + \underline{1} \text{ P} \rightarrow \underline{1} \text{ Rb}_3\text{P}$
- 27) $\underline{1} \text{ CH}_4 + \underline{2} \text{ O}_2 \rightarrow \underline{1} \text{ CO}_2 + \underline{2} \text{ H}_2\text{O}$
- 28) $\underline{2} \text{ Al(OH)}_3 + \underline{3} \text{ H}_2\text{SO}_4 \rightarrow \underline{1} \text{ Al}_2(\text{SO}_4)_3 + \underline{6} \text{ H}_2\text{O}$
- 29) $\underline{2} \text{ Na} + \underline{1} \text{ Cl}_2 \rightarrow \underline{2} \text{ NaCl}$
- 30) $\underline{16} \text{ Rb} + \underline{1} \text{ S}_8 \rightarrow \underline{8} \text{ Rb}_2\text{S}$
- 31) $\underline{2} \text{ H}_3\text{PO}_4 + \underline{3} \text{ Ca(OH)}_2 \rightarrow \underline{1} \text{ Ca}_3(\text{PO}_4)_2 + \underline{6} \text{ H}_2\text{O}$
- 32) $\underline{1} \text{ NH}_3 + \underline{1} \text{ HCl} \rightarrow \underline{1} \text{ NH}_4\text{Cl}$
- 33) $\underline{2} \text{ Li} + \underline{2} \text{ H}_2\text{O} \rightarrow \underline{2} \text{ LiOH} + \underline{1} \text{ H}_2$
- 34) $\underline{1} \text{ Ca}_3(\text{PO}_4)_2 + \underline{3} \text{ SiO}_2 + \underline{5} \text{ C} \rightarrow \underline{3} \text{ CaSiO}_3 + \underline{5} \text{ CO} + \underline{2} \text{ P}$
- 35) $\underline{4} \text{ NH}_3 + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ N}_2 + \underline{6} \text{ H}_2\text{O}$
- 36) $\underline{4} \text{ FeS}_2 + \underline{11} \text{ O}_2 \rightarrow \underline{2} \text{ Fe}_2\text{O}_3 + \underline{8} \text{ SO}_2$
- 37) $\underline{5} \text{ C} + \underline{2} \text{ SO}_2 \rightarrow \underline{1} \text{ CS}_2 + \underline{4} \text{ CO}$

Another Balancing Equations Worksheet Answers

Jicheng Xie



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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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Early Algebraization Jinfa Cai, Eric Knuth, 2011-02-24 In this volume the authors address the development of students algebraic thinking in the elementary and middle school grades from curricular cognitive and instructional perspectives The volume is also international in nature thus promoting a global dialogue on the topic of early Algebraization

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Try Out 7 Rao Indiraa Seshagiri,2008-09 Basics of Analytical Chemistry and Chemical Equilibria Brian M. Tissue,2013-06-06 Enables students to progressively build and apply new skills and knowledge Designed to be completed in one semester this text enables students to fully grasp and apply the core concepts of analytical chemistry and aqueous chemical equilibria Moreover the text enables readers to master common instrumental methods to perform a broad range of quantitative analyses Author Brian Tissue has written and structured the text so that readers progressively build their knowledge beginning with the most fundamental concepts and then continually applying these concepts as they advance to more sophisticated theories and applications Basics of Analytical Chemistry and Chemical Equilibria is clearly written and easy to follow with plenty of examples to help readers better understand both concepts and applications In addition there are several pedagogical features that enhance the learning experience including Emphasis on correct IUPAC terminology You Try It spreadsheets throughout the text challenging readers to apply their newfound knowledge and skills Online tutorials to build readers skills and assist them in working with the text s spreadsheets Links to analytical methods and instrument suppliers Figures illustrating principles of analytical chemistry and chemical equilibria End of chapter exercises Basics of Analytical Chemistry and Chemical Equilibria is written for undergraduate students who have completed a basic course in general chemistry In addition to chemistry students this text provides an essential foundation in analytical chemistry needed by students and practitioners in biochemistry environmental science chemical engineering materials science nutrition agriculture and the life sciences

Chemical Process Design and Simulation: Aspen Plus and Aspen Hysys Applications Juma Haydary,2019-01-23 A comprehensive and example oriented text for the study of chemical process design and simulation Chemical Process Design and Simulation is an accessible guide that offers information on the most important principles of chemical engineering design and includes illustrative examples of their application that uses simulation software A comprehensive and practical resource the text uses both Aspen Plus and Aspen Hysys simulation software The author describes the basic methodologies for computer aided design and offers a description of the basic steps of process simulation in Aspen Plus and Aspen Hysys The text reviews the design and simulation of individual simple unit operations that includes a mathematical model of each unit operation such as reactors separators and heat exchangers The author also explores the design of new plants and simulation of existing plants where conventional chemicals and material mixtures with measurable compositions are used In addition to aid in comprehension solutions to

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Chemical Process Design and Simulation is a practical and accessible guide to the chemical process design and simulation using proven software *Solutions Teacher Planning Pack Support Book 7* David Baker, 2005 The only AQA GCSE maths series to be exclusively endorsed and approved by AQA AQA Mathematics for GCSE blends print and electronic resources to provide you with complete reassurance that you have everything you need to deliver the revised 2006 GCSE Mathematics specification *Thermodynamics of the Earth and Planets* Alberto Patiño Douce, 2011-08-25 This textbook provides an intuitive yet mathematically rigorous introduction to the thermodynamics and thermal physics of planetary processes It demonstrates how the workings of planetary bodies can be understood in depth by reducing them to fundamental physics and chemistry The book is based on two courses taught by the author for many years at the University of Georgia It includes Guided Exercise boxes end of chapter problems worked solutions provided online and software boxes Maple code provided online As well as being an ideal textbook on planetary thermodynamics for advanced students in the Earth and planetary sciences it also provides an innovative and quantitative complement to more traditional courses in geological thermodynamics petrology chemical oceanography and planetary science In addition to its use as a textbook it is also of great interest to researchers looking for a one stop source of concepts and techniques that they can apply to their research problems **Me n Mine-Science-Term-1** Saraswati Experts, A text book on science *Ludwig's Applied Process Design for Chemical and Petrochemical Plants* A. Kayode Coker, 2011-08-30 This complete revision of Applied Process Design for Chemical and Petrochemical Plants Volume 1 builds upon Ernest E Ludwig s classic text to further enhance its use as a chemical engineering process design manual of methods and proven fundamentals This new edition includes important supplemental mechanical and related data nomographs and charts Also included within are improved techniques and fundamental methodologies to guide the engineer in designing process equipment and applying chemical processes to properly detailed equipment All three volumes of Applied Process Design for Chemical and Petrochemical Plants serve the practicing engineer by providing organized design procedures details on the equipment suitable for application selection and charts in readily usable form Process engineers designers and operators will find more chemical petrochemical plant design data in Volume 2 Third Edition which covers distillation and packed towers as well as material on azeotropes and ideal non ideal systems Volume 3 Third Edition which covers heat transfer refrigeration systems compression surge drums and

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mechanical vacuum systems process safety and pressure relieving devices metallurgy and corrosion and optimization of chemical process blending Some chapters review pressure relieving devices and provide case studies for process safety incidents which are well illustrated from US Chemical Safety Hazard Investigation Board www.csb.gov Finally this book contains a glossary of Petroleum and Petrochemical Terminologies and Physical and Chemical Characteristics of Major Hydrocarbons Provides improved design manual for methods and proven fundamentals of process design with related data and charts Covers complete range of basic day to day petrochemical operation topics Extensively revised with new material added on three phase separation metallurgy and corrosion Process safety management HAZOP and hazard analyses and optimization of chemical process blending Presents many examples using Honeywell UniSim Design software developed and executable computer programs and Excel spreadsheet programs Includes case studies of process safety incidents guidance for troubleshooting and checklists Includes Software of Conversion Table and 30 process data sheets in excel format

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