

Example Problem 10-3**Converting Mass to Moles**

A roll of copper wire has a mass of 545 g. How many moles of copper are in the roll? Use the atomic mass of copper given on the periodic table to apply a conversion factor to the mass given in the problem.

$$\text{moles of Cu} = \text{grams Cu} \times \frac{1 \text{ mol Cu}}{\text{grams Cu}}$$

$$\text{moles of Cu} = 545 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.546 \text{ g Cu}} = 8.58 \text{ mol Cu}$$

Note that the conversion factor is arranged so that g Cu cancels, leaving only mol Cu, the desired quantity.

Example Problem 10-4**Converting Moles to Mass**

Calculate the mass of 0.025 moles of calcium.

Use the molar mass of calcium to apply a conversion factor to the number of moles given in the problem. According to the periodic table, the atomic mass of calcium is 40.078 amu, so the molar mass of calcium is 40.078 g.

$$\text{mass of Ca} = \text{mol Ca} \times \frac{\text{grams Ca}}{1 \text{ mol Ca}}$$

$$\text{mass of Ca} = 0.025 \text{ mol Ca} \times \frac{40.078 \text{ g Ca}}{1 \text{ mol Ca}} = 1.00 \text{ g Ca}$$

The conversion factor is arranged so that mol Ca cancels, leaving g Ca, the desired quantity.

Practice Problems

- Calculate the mass of 0.88 mol selenium.
- A chemist needs 0.0788 mol selenium for a reaction. What mass of selenium should the chemist use?
- A sample of sulfur has a mass of 225 g. How many moles of sulfur are in the sample?

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Table of Contents Answer Key Chemistry The Mole Measuring Matter

1. Understanding the eBook Answer Key Chemistry The Mole Measuring Matter
 - The Rise of Digital Reading Answer Key Chemistry The Mole Measuring Matter
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemistry The Mole Measuring Matter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chemistry The Mole Measuring Matter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemistry The Mole Measuring Matter

- Personalized Recommendations
- Answer Key Chemistry The Mole Measuring Matter User Reviews and Ratings
- Answer Key Chemistry The Mole Measuring Matter and Bestseller Lists
- 5. Accessing Answer Key Chemistry The Mole Measuring Matter Free and Paid eBooks
 - Answer Key Chemistry The Mole Measuring Matter Public Domain eBooks
 - Answer Key Chemistry The Mole Measuring Matter eBook Subscription Services
 - Answer Key Chemistry The Mole Measuring Matter Budget-Friendly Options
- 6. Navigating Answer Key Chemistry The Mole Measuring Matter eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Chemistry The Mole Measuring Matter Compatibility with Devices
 - Answer Key Chemistry The Mole Measuring Matter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Chemistry The Mole Measuring Matter
 - Highlighting and Note-Taking Answer Key Chemistry The Mole Measuring Matter
 - Interactive Elements Answer Key Chemistry The Mole Measuring Matter
- 8. Staying Engaged with Answer Key Chemistry The Mole Measuring Matter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Chemistry The Mole Measuring Matter
- 9. Balancing eBooks and Physical Books Answer Key Chemistry The Mole Measuring Matter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chemistry The Mole Measuring Matter
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Chemistry The Mole Measuring Matter
 - Setting Reading Goals Answer Key Chemistry The Mole Measuring Matter
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Chemistry The Mole Measuring Matter

- Fact-Checking eBook Content of Answer Key Chemistry The Mole Measuring Matter
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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