

Practicing the Mole - - Even Problems Answer Key

Calculate the mass in grams of each of the following:

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol} \mid 26.98 \text{ g}}{1 \text{ mol}} = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I₂)

$$\frac{7.00 \text{ mol} \mid 253.8 \text{ g}}{1 \text{ mol}} = 1,780 \text{ g I}_2$$

4. 2.00×10^2 moles of chlorine (Cl₂)

$$\frac{200. \text{ mol} \mid 70.90 \text{ g}}{1 \text{ mol}} = 14,200 \text{ g Cl}_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol} \mid 55.85 \text{ g}}{1 \text{ mol}} = 514 \text{ g Fe}$$

Calculate the number of moles of atoms in each of the following:

10. 281 g of iron

$$\frac{281 \text{ g} \mid 1 \text{ mol}}{55.85 \text{ g}} = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g} \mid 1 \text{ mol}}{30.97 \text{ g}} = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g} \mid 1 \text{ mol}}{24.31 \text{ g}} = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g} \mid 1 \text{ mol}}{207.2 \text{ g}} = 4.995 \text{ mol Pb}$$

Calculate the number of atoms in each of the following:

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 6.02 \times 10^{24} \text{ atoms Ti}$$

Answer Key To Practtice Problems In Chemistry

Arshad Iqbal



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H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

Chemistry Workbook For Dummies Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence

Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading

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