Chemistry Student Edition - Basic Answer Key

Chapter 19: Chemical Equilibrium

The Nature of Chemical Equilibrium

Questions

- At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
- 2. What does the law of mass action say about any reaction at equilibrium?
- Why are solids and liquids left out of the equilibrium constant expression?
- Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
- Construct balanced chemical equations and write the equilibrium constant expression (either K_a or K_a) for each of the following processes:
 - a) Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
 - b) Aluminum oxide decomposes to form aluminum metal and oxygen gas.
 - Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
- Give an example of a chemical equilibrium that exists in our daily lives.
- Write a K_{ee} expression for each of the following processes.
 - Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
 - Hydrogen gas (H₂) and solid iodine (I₂) are combined to form hydrogen iodide gas (HI).

Answers

5.

- 1. At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction
- The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
- Solids and liquids have essentially constant densities, which means that writing them is not very useful.
- 4. Partial pressures are related to concentrations by a factor of RT.

$$\begin{split} &2\mathrm{Na}_{00}+\Diamond_{200} \stackrel{\longleftarrow}{\longleftarrow} \mathrm{\acute{E}}\ 2\mathrm{Na}_{2}\Diamond_{00}\\ &\mathrm{K}_{c}=\frac{[\mathrm{Na}_{2}\varnothing_{00}]^{2}}{[\mathrm{Na}_{00}]^{2}[\Diamond_{200}]} \end{split}$$

a) heterogeneous equilibria

$$2Al_2 \bigcirc_{300} \longrightarrow 4Al_{00} + 3\bigcirc_{300}$$

$$K_c = \frac{[AI'_{00}]^4 [\bigcirc_{200}]^3}{[AI_2 \varnothing_{80}]^2}$$

b) heterogeneous equilibria

Answer Key Ch 19 Chemistry

Arshad Iqbal

Answer Key Ch 19 Chemistry:

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nanomaterials display their functional property based on their nano and micro structure characterization is very important Topics found in Volume 2 include Sol gel technology is a versatile technology making it possible to produce a wide variety of materials and to provide existing substances with novel properties. This technology was applied to producing novel materials for example organic inorganic hybrids which are quite difficult to make by other fabricating techniques and it was also applied to producing materials based on high temperature superconducting oxides Applications of Sol Gel Technology Volume 3 will cover applications such as Chemical Thermodynamics for Industry Trevor M Letcher, 2007-10-31 Chemical Thermodynamics for Industry presents the latest developments in applied thermodynamics and highlights the role of thermodynamics in the chemical industry Written by leading experts in the field Chemical Thermodynamics for Industry covers the latest developments in traditional areas such as calorimetry microcalorimetry transport properties crystallization adsorption electrolyte systems and transport fuels It highlights newly established areas such as multiphase modeling reactive distillation non equilibrium thermodynamics and spectro calorimetry. It also explores new ways of treating old technologies as well as new and potentially important areas such as ionic liquids new materials ab initia quantum chemistry nano particles polymer recycling clathrates and the economic value of applied thermodynamics. This book is aimed not only at those working in a specific area of chemical thermodynamics but also at the general chemist the prospective researcher and those involved in funding chemical research Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi,2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

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