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Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-cl" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

1	H ⁺																	
2	Li ⁺	Be ²⁺									N ³⁻	O ²⁻	F ¹⁻					
3	Na ⁺	Mg ²⁺	Transition elements						Al ³⁺		P ³⁻	S ²⁻	Cl ¹⁻					
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺									Br ¹⁻			
5	Rb ⁺	Sr ²⁺			Ag ¹⁺									Sn ²⁺ Sn ⁴⁺		I ¹⁻		
6		Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺								Pb ²⁺ Pb ⁴⁺				
	Cations														Anions			

1. Consider the ions of potassium (K) and sulfur (S). Write the chemical formulas for the ionic compound involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
 K_2S
2. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

KS

K.S. K.S.

NaCl Sodium chloride

Zn₃P₂ Zinc phosphide

CaS Calcium sulfide

Al₂O₃ Aluminum oxide

Ag₂S Silver sulfide

SrCl₂ Strontium chloride

3. Circle the symbol for the metal in each of the compounds in Model 2. If doing electronically, you can just **BOLD** the symbol. Remember your PT!

Answer Key For Chapter Assessment Ionic Compounds

Neil D. Jespersen, Pamela Kerrigan



Answer Key For Chapter Assessment Ionic Compounds:

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