#### Reviewing Content

- The solvent is the substance in which the solute is dissolved.
- Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 44. solubility: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.

saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. unsaturated solution: a solution that contains less solute than a saturated solution at a given temperature and pressure. miscible: describes liquids that dissolve in each other. immiscible: describes liquids that are insoluble in each other.

- 45. Particles of solute crystallize.
- No: if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 47. 5.55 × 102 g AgNO<sub>3</sub>
- 48. Solubility increases with pressure.
- 49. a. 1.6×10<sup>-2</sup> g/L b. 4.7×10<sup>-2</sup> g/L
- Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- Molarity is the number of moles of solute dissolved in one liter of solution.
  - a. 1.3M KCI
  - b. 3.3 × 10<sup>-1</sup>M MgCl<sub>2</sub>
- 52.  $2.00 \times 10^{1}$  mL.
- a. 5.0×10<sup>-1</sup> mol NaCl, 29 g NaCl
  b. 1.0 mol KNO<sub>3</sub>, 1.0×10<sup>-2</sup> g KNO<sub>3</sub>
  c. 2.5×10<sup>2</sup> mol CaCl<sub>2</sub>, 2.8 g CaCl<sub>2</sub>
- 54. a. 2.3 × 10<sup>1</sup> g NaCl b. 2.0 g MgCl.
- 55. a. 16% (v/v) ethanol
  - b. 63.6% (v/v) isopropyl alcohol

- Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation. freezing-point depression, and vaporpressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution: relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
- 57. a. seawater
  - b. 1.5M KNO<sub>3</sub>
  - e. 0.100M MgCl<sub>2</sub>
- The effective molality of the Ca(NO<sub>3</sub>)<sub>2</sub> solution is 3m. The effective molality of the NaNO<sub>3</sub> solution is 2m.
- 59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
- The salt lowers the freezing point of the ice-water cooling mixture.
- 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
- 62. Add 27.0 g H2O to 32.0 g CH2OH.
- 63. a. 100.26°C
  - b. 101.54°C
- 64. a. -4.46°C
  - b. -2.2°C
- 65. a. -1.1°C
  - b. -0.74°C
  - e. −1.5°C

#### Understanding Concepts

 a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# **Answer Key For Chemistry Ch 16 Assessment**

**Angie Papple Johnston** 

#### **Answer Key For Chemistry Ch 16 Assessment:**

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Kaplan SAT Subject Test Chemistry 2015-2016 Kaplan Test Prep, 2015-03-03 Essential strategies practice and review to ace the SAT Subject Test Chemistry Getting into a top college has never been more difficult Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge Kaplan's SAT Subject Test Chemistry is the most up to date guide on the market with complete coverage of both the content review and strategies students need for success on test day Kaplan's SAT Subject Test Chemistry features A full length diagnostic test Full length practice tests Focused chapter summaries highlights and guizzes Detailed answer explanations Proven score raising strategies End of chapter guizzes Kaplan is serious about raising students scores we quarantee students will get a higher score **Chemistry for the Utterly Confused** John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject From formulas and lab techniques to the periodic table Chemistry for the Utterly Confused focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts This invaluable guide also teaches problem solving skills you need to master this imposing subject Whether you re in high school in college or simply brushing up on chemistry knowledge this fun easily accessible book will make understanding chemistry a breeze A Level Chemistry MCQ (Multiple Choice Questions) Arshad Igbal, 2019-06-18 The A Level Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF A Level Chemistry MCQ PDF Download Quiz Questions Chapter 1 28 Practice Tests with Answer Key IGCSE GCE Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs A Level Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests A Level Chemistry MCQ PDF book helps to practice test questions from exam prep notes The A Level Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs A Level Chemistry Multiple Choice Questions and Answers MCQs PDF

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