

Practice Sheet

Electron Dot (Lewis) Structures

A Lewis or Electron Dot Structure is a convenient representation of the valence electrons in an atom.

An electron dot structure for an atom is simply the symbol for the element, surrounded by a number of dots equal to the number of valence electrons.

Avoid a common mistake: the dots represent valence electrons only, so make sure you use only the number of dots corresponding to the number of valence electrons.

Examples:



- Draw Lewis dot structures for an atom of each of following elements:

1. K

3. Ar

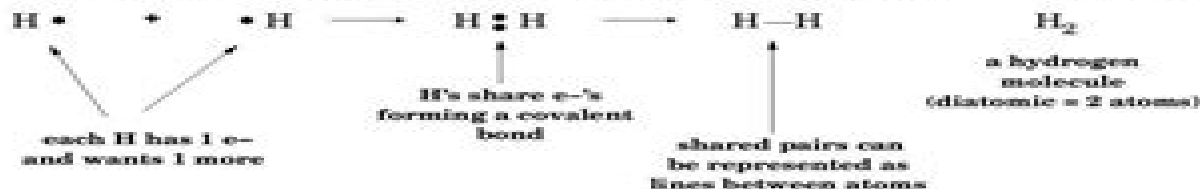
2. Si

4. As

Covalent Bonds

An atom can also achieve a noble gas configuration by sharing electrons with other atoms. When two non-metals combine, they typically share electrons in covalent bonds and form what are known as covalent compounds. We can draw Lewis Electron Dot Diagrams for covalent molecules to determine what type of bonds the molecules form.

For example, consider two hydrogen atoms. Each one has one valence electron. However, both hydrogens would like to have two electrons, to obtain the same configuration as [He]. To do this, each H atom shares an electron with the other, forming a covalent bond. In H_2 , each H atom has 2 electrons.



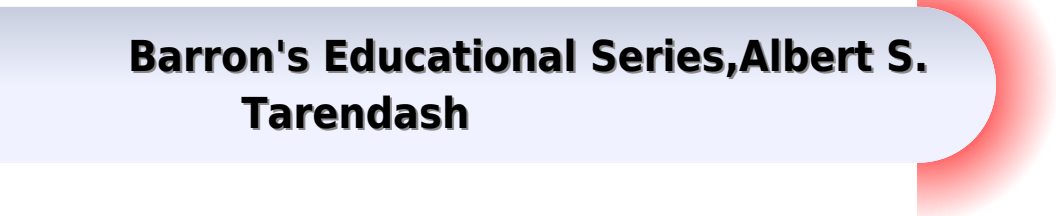
Here's another example, this time for chlorine. A Cl atom has 7 valence electrons, so it needs 1 more to achieve an octet (8 electrons).



Each chlorine now has 8 electrons, because you can count the shared ones:

Answer Key Of Lewis Dot Diagrams Chem

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Table of Contents Answer Key Of Lewis Dot Diagrams Chem

1. Understanding the eBook Answer Key Of Lewis Dot Diagrams Chem
 - The Rise of Digital Reading Answer Key Of Lewis Dot Diagrams Chem
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Of Lewis Dot Diagrams Chem
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Of Lewis Dot Diagrams Chem
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Of Lewis Dot Diagrams Chem
 - Personalized Recommendations
 - Answer Key Of Lewis Dot Diagrams Chem User Reviews and Ratings
 - Answer Key Of Lewis Dot Diagrams Chem and Bestseller Lists
5. Accessing Answer Key Of Lewis Dot Diagrams Chem Free and Paid eBooks
 - Answer Key Of Lewis Dot Diagrams Chem Public Domain eBooks
 - Answer Key Of Lewis Dot Diagrams Chem eBook Subscription Services
 - Answer Key Of Lewis Dot Diagrams Chem Budget-Friendly Options
6. Navigating Answer Key Of Lewis Dot Diagrams Chem eBook Formats

- ePub, PDF, MOBI, and More
 - Answer Key Of Lewis Dot Diagrams Chem Compatibility with Devices
 - Answer Key Of Lewis Dot Diagrams Chem Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Of Lewis Dot Diagrams Chem
 - Highlighting and Note-Taking Answer Key Of Lewis Dot Diagrams Chem
 - Interactive Elements Answer Key Of Lewis Dot Diagrams Chem
 8. Staying Engaged with Answer Key Of Lewis Dot Diagrams Chem
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Of Lewis Dot Diagrams Chem
 9. Balancing eBooks and Physical Books Answer Key Of Lewis Dot Diagrams Chem
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Of Lewis Dot Diagrams Chem
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Key Of Lewis Dot Diagrams Chem
 - Setting Reading Goals Answer Key Of Lewis Dot Diagrams Chem
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Key Of Lewis Dot Diagrams Chem
 - Fact-Checking eBook Content of Answer Key Of Lewis Dot Diagrams Chem
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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