

Analytical Chemistry MCQs: TYBSC SEM III

8. The solubility product K_{sp} of sparingly soluble Ag_2CrO_4 is 4×10^{-12} . The molar solubility of the salt is
- (a) 1.0×10^{-4} moles/lit
 - (b) 2.0×10^{-4} moles/lit
 - (c) 1.0×10^{-5} moles/lit
 - (d) 1.2×10^{-12} moles/lit
9. Solubility product of $AgCl$ is 1.44×10^{-4} moles/lit at 373K. The solubility of the $AgCl$ in boiling of water will be....
- (a) 0.72×10^{-4} moles/lit
 - (b) 0.72×10^{-2} moles/lit
 - (c) 1.20×10^{-2} moles/lit
 - (d) 1.20×10^{-4} moles/lit
10. Precipitation takes place when the ionic product....
- (a) Is zero
 - (b) Exceeds the solubility product
 - (c) Equal to solubility product
 - (d) Is less than the solubility product
11. The electromotive force is expressed in the unit of
- (a) Ohm
 - (b) Volt
 - (c) Ampere
 - (d) Coulomb
12. If the cell can provide electrical energy to an external system, it is called as
- (a) Galvanic cell
 - (b) Electrolytic cell
 - (c) Concentration cell
 - (d) Chemical cell
13. The current per unit area of electrode surface is known as....
- (a) Current efficiency
 - (b) Current yield
 - (c) Current density
 - (d) None of these

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Table of Contents Analytical Chemistry Mcqs With Answers

1. Understanding the eBook Analytical Chemistry Mcqs With Answers
 - The Rise of Digital Reading Analytical Chemistry Mcqs With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Analytical Chemistry Mcqs With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analytical Chemistry Mcqs With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analytical Chemistry Mcqs With Answers

- Personalized Recommendations
- Analytical Chemistry Mcqs With Answers User Reviews and Ratings
- Analytical Chemistry Mcqs With Answers and Bestseller Lists
- 5. Accessing Analytical Chemistry Mcqs With Answers Free and Paid eBooks
 - Analytical Chemistry Mcqs With Answers Public Domain eBooks
 - Analytical Chemistry Mcqs With Answers eBook Subscription Services
 - Analytical Chemistry Mcqs With Answers Budget-Friendly Options
- 6. Navigating Analytical Chemistry Mcqs With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Analytical Chemistry Mcqs With Answers Compatibility with Devices
 - Analytical Chemistry Mcqs With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analytical Chemistry Mcqs With Answers
 - Highlighting and Note-Taking Analytical Chemistry Mcqs With Answers
 - Interactive Elements Analytical Chemistry Mcqs With Answers
- 8. Staying Engaged with Analytical Chemistry Mcqs With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analytical Chemistry Mcqs With Answers
- 9. Balancing eBooks and Physical Books Analytical Chemistry Mcqs With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analytical Chemistry Mcqs With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analytical Chemistry Mcqs With Answers
 - Setting Reading Goals Analytical Chemistry Mcqs With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analytical Chemistry Mcqs With Answers

- Fact-Checking eBook Content of Analytical Chemistry Mcqs With Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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