GAS LAW WORKSHEET 1 KEY

 A sample of oxygen gas occupies a volume of 436.2 mL at 1.02 atm. If the temperature is held constant, what would the pressure of this gas be when the gas is compressed to 231.6 mL?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \text{ since } T_1 = T_2 \quad P_1V_1 = P_2V_2 \quad P_2 = P_1\frac{V_1}{V_2} = 1.02 \text{ atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm} \\ \left(\frac{436.2 \text{ mL}}{231.6 \text{ mL}}\right) = \textbf{1.92 atm}$$

To check with reasoning: decreased V leads to increased P so the volume ratio should be greater than one to yield the expected increased pressure.

 If a gas originally occupying 6.75 L at 19.21 °C and 762.5 torr is compressed to give a pressure of 1.26 atm at 26.35 °C, what would the new volume be?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \qquad V_2 = V_1 \left(\frac{T_2}{T_1}\right) \left(\frac{P_1}{P_2}\right) = 6.75 \text{ L} \left(\frac{299.50 \text{ K}}{292.36 \text{ K}}\right) \left(\frac{762.5 \text{ torr}}{1.26 \text{ atm}}\right) \left(\frac{1 \text{ atm}}{760 \text{ torr}}\right) = 5.51 \text{ L}$$

To check with reasoning: increased T leads to increased V, so the temperature ratio should be greater than one to yield the expected increase in volume. Increased P leads to decreased V, so the pressure ratio should be less than one to yield the expected decreased volume.

 Calculate the number of grams of hydrogen sulfide gas, H₂S, in a 3.6 L container at 32.6 °C and 712 mmHg.

$$PV = \frac{g}{M}RT - g = \frac{PVM}{RT} = \frac{712 \text{ mmHg (3.6 L) } 34.080 \frac{g}{\text{mol}}}{0.082058 \frac{L \text{ atm}}{K \text{ mol}}} \left(\frac{1 \text{ atm}}{760 \text{ mmHg}}\right) = 4.6 \text{ g H}_2S$$

Calculate the density of 0.625 g of carbon dioxide at 26.32 °C and 1.03 atm.

$$PV = \frac{g}{M}RT - \frac{g}{V} = \frac{PM}{RT} = \frac{1.03 \text{ atm} \left(44.010 \frac{g}{\text{mol}}\right)}{0.082058 \frac{L \text{ atm}}{K \text{ mol}} \left(299.47 \text{ K}\right)} = 1.84 \text{ g/L CO}_2$$

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Didier Musso

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