

8 • Electron Configurations & Periodicity

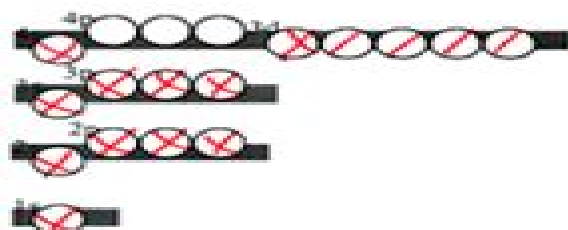
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
Ar $[Ne] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
Li $[He] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Answer Key Electron Configurations Of Atoms And Ions Lab

M Mosston



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Table of Contents Answer Key Electron Configurations Of Atoms And Ions Lab

1. Understanding the eBook Answer Key Electron Configurations Of Atoms And Ions Lab
 - The Rise of Digital Reading Answer Key Electron Configurations Of Atoms And Ions Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Electron Configurations Of Atoms And Ions Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Electron Configurations Of Atoms And Ions Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Electron Configurations Of Atoms And Ions Lab
 - Personalized Recommendations
 - Answer Key Electron Configurations Of Atoms And Ions Lab User Reviews and Ratings
 - Answer Key Electron Configurations Of Atoms And Ions Lab and Bestseller Lists
5. Accessing Answer Key Electron Configurations Of Atoms And Ions Lab Free and Paid eBooks
 - Answer Key Electron Configurations Of Atoms And Ions Lab Public Domain eBooks
 - Answer Key Electron Configurations Of Atoms And Ions Lab eBook Subscription Services
 - Answer Key Electron Configurations Of Atoms And Ions Lab Budget-Friendly Options

6. Navigating Answer Key Electron Configurations Of Atoms And Ions Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Electron Configurations Of Atoms And Ions Lab Compatibility with Devices
 - Answer Key Electron Configurations Of Atoms And Ions Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Highlighting and Note-Taking Answer Key Electron Configurations Of Atoms And Ions Lab
 - Interactive Elements Answer Key Electron Configurations Of Atoms And Ions Lab
8. Staying Engaged with Answer Key Electron Configurations Of Atoms And Ions Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Electron Configurations Of Atoms And Ions Lab
9. Balancing eBooks and Physical Books Answer Key Electron Configurations Of Atoms And Ions Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Electron Configurations Of Atoms And Ions Lab
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Electron Configurations Of Atoms And Ions Lab
 - Setting Reading Goals Answer Key Electron Configurations Of Atoms And Ions Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Fact-Checking eBook Content of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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