

Student Exploration: Calorimetry Lab

(Note: Gizmos can be done even on your phones! Contact Goff or Molik if you need your log-in information)

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



1. On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

- A. What was the **Final temperature** of the copper and the water? 35 °C _____
- B. How much did the temperature of the copper change? It went down 65 °C _____
- C. How much did the temperature of the water change? _____ it went up 5 °C _____

2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Activity A: Heat transfer	<u>Get the Gizmo ready:</u> • Click Reset (↺).	
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Question: What factors determine how heat energy transfers between objects?

1. Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

- A. How do you think increasing the water's mass would affect the final temperature?

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SA Adler



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Table of Contents Answer Key Explore Learning Calorimetry Lab

1. Understanding the eBook Answer Key Explore Learning Calorimetry Lab
 - The Rise of Digital Reading Answer Key Explore Learning Calorimetry Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Explore Learning Calorimetry Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Explore Learning Calorimetry Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Explore Learning Calorimetry Lab

- Personalized Recommendations
- Answer Key Explore Learning Calorimetry Lab User Reviews and Ratings
- Answer Key Explore Learning Calorimetry Lab and Bestseller Lists
- 5. Accessing Answer Key Explore Learning Calorimetry Lab Free and Paid eBooks
 - Answer Key Explore Learning Calorimetry Lab Public Domain eBooks
 - Answer Key Explore Learning Calorimetry Lab eBook Subscription Services
 - Answer Key Explore Learning Calorimetry Lab Budget-Friendly Options
- 6. Navigating Answer Key Explore Learning Calorimetry Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Explore Learning Calorimetry Lab Compatibility with Devices
 - Answer Key Explore Learning Calorimetry Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Explore Learning Calorimetry Lab
 - Highlighting and Note-Taking Answer Key Explore Learning Calorimetry Lab
 - Interactive Elements Answer Key Explore Learning Calorimetry Lab
- 8. Staying Engaged with Answer Key Explore Learning Calorimetry Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Explore Learning Calorimetry Lab
- 9. Balancing eBooks and Physical Books Answer Key Explore Learning Calorimetry Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Explore Learning Calorimetry Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Explore Learning Calorimetry Lab
 - Setting Reading Goals Answer Key Explore Learning Calorimetry Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Explore Learning Calorimetry Lab

- Fact-Checking eBook Content of Answer Key Explore Learning Calorimetry Lab
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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