

# Chemistry Worksheets Class 11 on Chapter 4 Chemical Bonding and Molecular Structure - Set 3

Q-1: Which of the	following	possesses an	expanded	octet?
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- a) IF,
- b) SF.
- c) BrF<sub>n</sub>
- d) All of the above

Q-2: Strongly electronegative element B contrasts with strongly electropositive element A. Both are divalent. The resulting compound would be

- a) A-B
- b) A2\*B2-
- c) A2-B2"
- d) A-B

Q-3: Which of the following compounds shows the highest lattice energy?

- a) CsCl
- b) KCI
- c) RbCl
- d) NaCl

Q-4: Which one of the following molecules is formed by p-p overlapping?

- a)  $I_2$
- b) HBr
- c) H<sub>2</sub>S
- d) PH<sub>a</sub>

Q-5: Which of the following H-bonds has the greatest impact on a molecule's physical characteristics?

- a) Intramolecular H-bond
- b) Intermolecular H-bond
- c) Both of the above
- d) All H-bonds

Q-6: State the crucial conditions that must be met for a molecule to undergo hybridisation.

Q-7: Describe why CH<sub>4</sub> has a tetrahedral geometry rather than a square planar geometry with a carbon atom in the centre and four H atoms at each corner.

# **Answer Key For Holt Science Chemical Bonding**

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