



Directions (1-10): For each statement or question, choose the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Science/Chemistry.

1. The oxidation number of nitrogen in N_2 is

- (1) +1 (3) +3
☒ (2) 0 (4) -3

2. What is the oxidation number of Pt in K_2PtCl_6 ?

- (1) -2 (3) -4
(2) +2 ☒ (4) +4

3. In which substance does phosphorus have a +3 oxidation state?

- (1) P_4O_{10} (3) $Ca_3(PO_4)_2$
(2) PCl_5 ☒ (4) KH_2PO_4

4. What is the oxidation state of sulfur in H_2SO_4 ?

- (1) 0 ☒ (3) +6
(2) -2 (4) +4

5. In which substance does sulfur have a negative oxidation number?

- ☒ (1) Na_2S (3) S_8
(2) $CaSO_4$ (4) SO_2

6. In which compound is the oxidation number of oxygen -1?

- (1) CO (3) H_2O
(2) CO_2 ☒ (4) H_2O_2

7. Oxygen has an oxidation number of -2 in

- (1) O_2 (3) Na_2O_2
☒ (2) NO_2 (4) OF_2

8. Oxygen will have a positive oxidation number when combined with

- ☒ (1) fluorine (3) bromine
(2) chlorine (4) iodine

9. In which compound does hydrogen have an oxidation number of -1?

- (1) NH_3 (3) HCl
☒ (2) KH (4) H_2O

10. Which equation represents an oxidation-reduction reaction?

- (1) $HCl + KOH \rightarrow KCl + H_2O$
☒ (2) $4 HCl + MnO_2 \rightarrow MnCl_2 + H_2O + Cl_2$
(3) $2 HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$
(4) $2 HCl + FeS \rightarrow FeCl_2 + H_2S$

11. Which equation represents an oxidation-reduction reaction?

- ☒ (1) $Zn + 2 HCl \rightarrow ZnCl_2 + H_2$
(2) $Zn(OH)_2 + 2 HCl \rightarrow ZnCl_2 + 2 H_2O$
(3) $H_2O + NH_3 \rightarrow NH_4^+ + OH^-$
(4) $H_2O + H_2O \rightarrow H_3O^+ + OH^-$

12. Which statement correctly describes a redox reaction?

- ☒ (1) Oxidation and reduction occur simultaneously.
(2) Oxidation occurs before reduction.
(3) Oxidation occurs after reduction.
(4) Oxidation occurs, but reduction does not.

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Robert R. Redfield



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Table of Contents Answer Key For January 2015 Chemistry Regents

1. Understanding the eBook Answer Key For January 2015 Chemistry Regents
 - The Rise of Digital Reading Answer Key For January 2015 Chemistry Regents
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For January 2015 Chemistry Regents
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For January 2015 Chemistry Regents
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For January 2015 Chemistry Regents
 - Personalized Recommendations
 - Answer Key For January 2015 Chemistry Regents User Reviews and Ratings
 - Answer Key For January 2015 Chemistry Regents and Bestseller Lists
5. Accessing Answer Key For January 2015 Chemistry Regents Free and Paid eBooks
 - Answer Key For January 2015 Chemistry Regents Public Domain eBooks
 - Answer Key For January 2015 Chemistry Regents eBook Subscription Services
 - Answer Key For January 2015 Chemistry Regents Budget-Friendly Options
6. Navigating Answer Key For January 2015 Chemistry Regents eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For January 2015 Chemistry Regents Compatibility with Devices
 - Answer Key For January 2015 Chemistry Regents Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For January 2015 Chemistry Regents
 - Highlighting and Note-Taking Answer Key For January 2015 Chemistry Regents
 - Interactive Elements Answer Key For January 2015 Chemistry Regents
8. Staying Engaged with Answer Key For January 2015 Chemistry Regents

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Key For January 2015 Chemistry Regents
- 9. Balancing eBooks and Physical Books Answer Key For January 2015 Chemistry Regents
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For January 2015 Chemistry Regents
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key For January 2015 Chemistry Regents
 - Setting Reading Goals Answer Key For January 2015 Chemistry Regents
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For January 2015 Chemistry Regents
 - Fact-Checking eBook Content of Answer Key For January 2015 Chemistry Regents
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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