

Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

- 3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Answer Key For Stoichiometry Test

Julia Schneider



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Table of Contents Answer Key For Stoichiometry Test

1. Understanding the eBook Answer Key For Stoichiometry Test
 - The Rise of Digital Reading Answer Key For Stoichiometry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Stoichiometry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Stoichiometry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Stoichiometry Test
 - Personalized Recommendations
 - Answer Key For Stoichiometry Test User Reviews and Ratings
 - Answer Key For Stoichiometry Test and Bestseller Lists
5. Accessing Answer Key For Stoichiometry Test Free and Paid eBooks
 - Answer Key For Stoichiometry Test Public Domain eBooks
 - Answer Key For Stoichiometry Test eBook Subscription Services
 - Answer Key For Stoichiometry Test Budget-Friendly Options
6. Navigating Answer Key For Stoichiometry Test eBook Formats

- ePub, PDF, MOBI, and More
- Answer Key For Stoichiometry Test Compatibility with Devices
- Answer Key For Stoichiometry Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Stoichiometry Test
 - Highlighting and Note-Taking Answer Key For Stoichiometry Test
 - Interactive Elements Answer Key For Stoichiometry Test
- 8. Staying Engaged with Answer Key For Stoichiometry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Stoichiometry Test
- 9. Balancing eBooks and Physical Books Answer Key For Stoichiometry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Stoichiometry Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key For Stoichiometry Test
 - Setting Reading Goals Answer Key For Stoichiometry Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For Stoichiometry Test
 - Fact-Checking eBook Content of Answer Key For Stoichiometry Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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