

Limiting Reagent Worksheet #2

1. Consider the reaction



- a) 80.0 grams of diiodine pentoxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO.

Determine the mass of iodine I_2 , which could be produced?

- b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
i) what mass of iodine was produced?
ii) what percentage yield of iodine was produced.

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are reacted,

- a) Which chemical is the limiting reactant?
b) How many grams of ZnS will be formed?
c) How many grams of the excess reactant will remain after the reaction is over?
3. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen?
What mass is in excess? What mass of MgO is formed?

4. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S?

5. When MoO_3 and Zn are heated together they react



What mass of ZnO is formed when 20.0 grams of MoO_3 is reacted with 10.0 grams of Zn?

6. Silver nitrate, AgNO_3 , reacts with ferric chloride, FeCl_3 , to give silver chloride, AgCl , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$. In a particular experiment, it was planned to mix a solution containing 25.0 g of AgNO_3 with another solution containing 45.0 grams of FeCl_3 .

- a) Write the chemical equation for the reaction.
b) Which reactant is the limiting reactant?
c) What is the maximum number of moles of AgCl that could be obtained from this mixture?
d) What is the maximum number of grams of AgCl that could be obtained?
e) How many grams of the reactant in excess will remain after the reaction is over?

7. Solid calcium carbonate, CaCO_3 , is able to remove sulphur dioxide from waste gases by the reaction (balanced as written):



In a particular experiment, 255 g of CaCO_3 was exposed to 135 g of SO_2 in the presence of an excess amount of the other chemicals required for the reaction.

- a) What is the theoretical yield of CaSO_3 ?
b) If only 198 g of CaSO_3 was isolated from the products, what was the percentage yield of CaSO_3 in this experiment?

Answer Key To Limiting Reagent

Disha Experts



Answer Key To Limiting Reagent:

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1. Understanding the eBook Answer Key To Limiting Reagent
 - The Rise of Digital Reading Answer Key To Limiting Reagent
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Limiting Reagent
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Limiting Reagent
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Limiting Reagent
 - Personalized Recommendations
 - Answer Key To Limiting Reagent User Reviews and Ratings
 - Answer Key To Limiting Reagent and Bestseller Lists
5. Accessing Answer Key To Limiting Reagent Free and Paid eBooks
 - Answer Key To Limiting Reagent Public Domain eBooks
 - Answer Key To Limiting Reagent eBook Subscription Services
 - Answer Key To Limiting Reagent Budget-Friendly Options
6. Navigating Answer Key To Limiting Reagent eBook Formats

- ePub, PDF, MOBI, and More
 - Answer Key To Limiting Reagent Compatibility with Devices
 - Answer Key To Limiting Reagent Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Limiting Reagent
 - Highlighting and Note-Taking Answer Key To Limiting Reagent
 - Interactive Elements Answer Key To Limiting Reagent
 8. Staying Engaged with Answer Key To Limiting Reagent
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Limiting Reagent
 9. Balancing eBooks and Physical Books Answer Key To Limiting Reagent
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Limiting Reagent
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Key To Limiting Reagent
 - Setting Reading Goals Answer Key To Limiting Reagent
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Key To Limiting Reagent
 - Fact-Checking eBook Content of Answer Key To Limiting Reagent
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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