

NUCLEAR CHEMISTRY WORKSHEET

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Most nuclear reactions involve changes in the number of protons and/or neutrons in the nucleus of an atom. These changes are called *transmutations* because an atom of one element is changed into an atom of a different element. During these changes, nuclear particles may be absorbed and/or emitted along with a release of energy.

There are four types of nuclear reactions:

- 1) *natural radioactive decay*: transmutations occur by natural radioactivity (the ability of a nucleus to emit a nuclear particle and energy without external stimulation);
- 2) *artificial transmutation*: transmutations occur by nuclear disintegration caused by external stimulation as a scientist bombards a nucleus with a particle (the addition of another nuclear particle makes the nucleus unstable);
- 3) *fission*: certain nuclei having a large mass are bombarded with special particles that cause the nuclei to split into two nuclei each having a smaller mass;
- 4) *fusion*: nuclei of light elements are combined to form heavier nuclei.

Nuclear reactions are written using symbols in the A_ZX notation. The symbols for the major nuclear particles involved in nuclear reactions are given below.

Alpha particle	${}^4_2\text{He}$ (an alpha particle is a helium nucleus)
Proton	${}^1_1\text{H}$ (the most common hydrogen nucleus is a proton)
Neutron	${}^1_0\text{n}$
Electron	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$ (also called a beta particle)
Positron	${}^0_1\text{e}$ or ${}^0_1\beta$
Gamma ray	${}^0_0\gamma$

The table below contains information about nuclear reactions involving the emission of radiation.

Type of Reaction	NUCLEAR CHANGES		
	Radiation (Particle)	Effect on the Atomic Number	Effect on the Atomic Mass
Alpha emission (α)	${}^4_2\text{He}$	decrease by 2	decrease by 4
Beta emission (β)	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	increase by 1	no change
Positron emission (β^+)	${}^0_1\text{e}$ or ${}^0_1\beta$	decrease by 1	no change
Gamma emission	${}^0_0\gamma$	no change	no change
Electron capture (EC)	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	decrease by 1	no change

Answer Key To Nuclear Chemistry

Mark Steyvers



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Table of Contents Answer Key To Nuclear Chemistry

1. Understanding the eBook Answer Key To Nuclear Chemistry
 - The Rise of Digital Reading Answer Key To Nuclear Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Nuclear Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Nuclear Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Nuclear Chemistry
 - Personalized Recommendations
 - Answer Key To Nuclear Chemistry User Reviews and Ratings
 - Answer Key To Nuclear Chemistry and Bestseller Lists
5. Accessing Answer Key To Nuclear Chemistry Free and Paid eBooks
 - Answer Key To Nuclear Chemistry Public Domain eBooks
 - Answer Key To Nuclear Chemistry eBook Subscription Services
 - Answer Key To Nuclear Chemistry Budget-Friendly Options

6. Navigating Answer Key To Nuclear Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To Nuclear Chemistry Compatibility with Devices
 - Answer Key To Nuclear Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Nuclear Chemistry
 - Highlighting and Note-Taking Answer Key To Nuclear Chemistry
 - Interactive Elements Answer Key To Nuclear Chemistry
8. Staying Engaged with Answer Key To Nuclear Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Nuclear Chemistry
9. Balancing eBooks and Physical Books Answer Key To Nuclear Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Nuclear Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Nuclear Chemistry
 - Setting Reading Goals Answer Key To Nuclear Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Nuclear Chemistry
 - Fact-Checking eBook Content of Answer Key To Nuclear Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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