



Allen Bradley Flex 40 Manual

Robert R. Redfield



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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acids bases and enzymes many acids and bases in living things provide the ph that enzymes need enzymes are biological catalysts that must work effectively for biochemical reactions to occur most enzymes can do their job only at a certain level of acidity cells secrete acids and bases to maintain the proper ph for enzymes to do their work

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acid base titration difference between acid and base acids and bases definition an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste

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arrhenius s definition of acids and bases the earliest definition of acids and bases is arrhenius s definition which states that

an acid is a substance that forms hydrogen ions H^+ when dissolved in water and a base is a substance that forms hydroxide ions OH^- when dissolved in water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is

$$\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$$

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