

American Chemical Society (A.C.S.) General Chemistry Exam

75 multiple choice questions, 110 minutes

Bring a scanner, calculator and pencil(s)

- I. States of Matter**
Ideal Gas Law: $PV = nRT$, $PM = dRT$, deviations from ideal (high P, low T)
Ionic Solids: sc, fcc, bcc
Phase Diagrams: Pressure versus Temperature
Intermolecular Forces: ion-dipole, hydrogen-bonding, etc.
- II. Stoichiometry and Thermochemistry**
Hess' Law & Enthalpy: "standard states" have $\Delta H = 0$
Percent Composition: %C, %H, etc.
Limiting Reactant(s)
Stoichiometry
- III. Atomic Structure and Periodicity**
Oxidation States of "Main group" elements (Li, Be, B, etc.)
Oxidation Numbers
Atomic orbital configurations & relation to groups on periodic table (ns^2np^2 = halogens)
Periodic Trends: Electronegativity, atomic radii
Nuclear chemistry: mass number, atomic number
- IV. Molecular Structure**
Modes of bonding: ionic, covalent (both polar and nonpolar)
Lewis Dot, VSEPR, polarity, bond angles, resonance structures
- V. Solutions**
Molarity: mol solute / L solution, also calculating molarity, $C_1V_1 = C_2V_2$
Net ionic equations
Percent dissociation in water (electrolyte strength)
Molality: mol solute / kg solvent
Freezing Point Depression $\Delta T = k_{fp}$
- VI. Acids, Bases and Ionic Equilibria**
Know the "inert" acid/base ions: Cl^- , NO_3^- , Na^+ , etc.
Know general reactions for $WA + SB$, $WB + SA$, $SB + SA$
Equivalence point, finding unknown concentrations
Buffer: what makes a buffer, what a buffer is, etc.
pH: calculate pH of $WB + SA$, $SB + WA$, pure WB , pure WA
- VII. Chemical Equilibria - Molecular**
General equilibrium expressions: know how to find K
Precipitation reactions
Effect of heat upon endothermic / exothermic reactions
- VIII. Kinetics**
Rate of disappearance, rate of formation, Catalysis in mechanisms
Rate law and half-life from experimental data; Energy of activation
- IX. Thermo-dynamics**
Enthalpy, entropy, free energy
Relation of free energy to equilibrium constants
- X. Electrochemistry and Redox**
E values for electrochemical cells; E values for electrolytic cells
Balancing acidic redox reactions
Best reducing agents, oxidizing agents, definitions of oxidized and reduced
- XI. Descriptive Chemistry**
Nuclear chemistry: particle(s) emitted which result in increase of atomic number
Elemental reactivity: with water, oxygen, bases, acid, etc.
Organic chemistry: know alkenes, alkanes, etc.; also nomenclature
Coordination chemistry: coordination numbers around central metal
Ionization energy periodic trends
General nomenclature
- XII. Laboratory**
Effect of improper drying on mass measurements
Methods for finding $[Cl^-]$

American Chemical Society A C S General Chemistry Exam

Michael Rajnik



American Chemical Society A C S General Chemistry Exam:

Reviewing **American Chemical Society A C S General Chemistry Exam**: Unlocking the Spellbinding Force of Linguistics

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