9) A 250mL solution of 0.5M NaF has 50mL of a 1M BaF- solution added to it. What is the concentration of [F]: in the new solution? Assume that volume is additive.

From (8)
$$[F] = 0.125 \text{ mol} [F] = 0.125 \text{ mol} + 0.1 \text{ mol} = 0.225 \text{ mol} [F] = 0.125 \text{ mol} + 0.1 \text{ mol} = 0.225 \text{ mol} [F] = 0.125 \text{ mol} = 0.225 \text{ mol} = 0.75 \text{ mol} = 0.225 \text{ mol} = 0.75 \text{ mol} = 0.225 \text{ mol} = 0.75 \text{ mol} = 0.30 \text{ mol} = 0.31 \text{ mol} = 0.75 \text{ mol} =$$

10) How many grams of NaHCO, (molar mass = 84.01 g/mol) are needs 350mt, solution of 0.75M NaHCO+?

1th Color date the amount of distilled water that is executed to dilute a 31 solution of 5M HBr to a concentration of 15M HBr3

$$mal = 3L \times 5m$$

V= 15mel/15MBR

12) What volume of 8M H+SQ, is needed to prepare a 100mL solution of 0.5M HUSOU?

O.JLXO.5MH2SOL = 0.005md H_50, neede

13) The maximum amount of sait that can be contained in water is approximately 26.3% by weight. Convert this into concentration with units of molarity, Info: 1g of waster a 1 ml of water

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Huangqi Zhang

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