

Unit 3 Review - Atomic Structure & Periodic Table

Part 1: Answer the following questions.

1. a. Which scientist created the first modern atomic theory? John Dalton
 b. What was his theory?
 1. All matter is made of extremely small particles called atoms.
 2. All atoms of a given element are identical and atoms of a specific element are different from any other element
 3. Atoms cannot be created, divided into smaller particles or destroyed
 4. Different atoms combine in simple whole number ratios to form compounds
 5. In a chemical reaction, atoms are separated, combined or rearranged.
- c. What was his model of the atom called? Billiard Ball
- d. What was proven wrong and why?
 1. All atoms of a given element are identical – wrong b/c of the existence of isotopes
 2. Atoms cannot be divided into smaller particles – wrong b/c of the existence of electrons, protons and neutrons
2. a. Which scientist did the gold foil experiment? Rutherford
 b. What did he conclude from the gold foil experiment?
 1. Atoms are made up of mostly empty space
 2. Atoms have a small positive center call the nucleus
3. a. Who discovered the electron? J. J. Thomson
 b. What was the name of his experiment? Cathode Ray Tube Experiment
 c. What was his model of the atom called? Plum Pudding Model
4. What did Mendeleev do?
Arranged the periodic table in order of increasing atomic mass
5. a. Who arranged the **modern** periodic table? Moseley
 b. How is the modern periodic table arranged? Arranged in order of Atomic Number
6. a. What are the three subatomic particles? Give the symbols, charge, mass and location of each.

Particles	Symbol	Charge	Mass	Location
Proton	p^+	Positive	1 amu	Nucleus
Neutron	n^0	Neutral	1 amu	Nucleus
Electron	e^-	Negative	1/2000 amu	Electron Cloud

7. What is special about the proton? The proton IDENTIFIES THE ELEMENT
8. Fill in the blanks in the following:
 - a. Mass # = $p^+ + n^0$
 - b. Atomic # = # of protons
9. What is the difference between Average Atomic mass and mass #? (Average) Atomic Mass is found on the periodic table and is the weighted average of the isotopes. The Mass Number is the mass of one element and must be given or calculated by adding the protons and neutron of that element.
10. Which of the following are isotopes of the same element? How do you know? A, and D. They have the same Atomic Number (# of Protons)
 - a. $^{12}_5X$
 - b. $^{15}_5X$
 - c. $^{12}_6X$
 - d. $^{15}_5X$
11. Calculate the average atomic mass for Boron if it has two naturally occurring isotopes: boron -10 (abundance of 19.8%) and boron - 11 (abundance of 80.2%).

$$\left(10 \times \frac{19.8}{100} \right) + \left(11 \times \frac{80.2}{100} \right) = 10.802 \text{ amu}$$
12. Identify an element "X" by calculating the average atomic mass. X = 6 (abundance of 7.5%) and X= 7 (abundance 92.5%).

$$\left(6 \times \frac{7.5}{100} \right) + \left(7 \times \frac{92.5}{100} \right) = 6.925 \text{ amu} \quad \text{Which is Lithium}$$

Answer Key Unit 3 Review Electrons

Lei Shi



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Mosby's Radiation Therapy Study Guide and Exam Review Leia Levy, 2025-08-21 Reinforce your understanding of radiation therapy and prepare for the Registry exam Mosby's Radiation Therapy Study Guide and Exam Review Second Edition is both a study companion for Washington and Leaver's Principles and Practice of Radiation Therapy and a superior review for the ARRT Radiation Therapy Certification Exam This completely updated edition reflects the latest exam specifications and features an easy to read format that presents information in concise bullets and tables More than 2 000 total multiple choice questions in Registry format provide a realistic testing experience to prepare you for the real exam NEW Quality control procedures and guidelines for linear accelerators NEW Fractionation and protraction considerations UPDATED Content reflects the latest ARRT Radiation Therapy Certification Exam Specifications UPDATED Radiation Protection and Safety and Overview of Cancer Imaging and Management Modalities chapters offer the most current information in these key areas EXPANDED Charge capture and record keeping content enhances the Oncology Patient Care chapter EXPANDED Additional cross sectional images provide anatomy review and reinforce treatment planning concepts More than 2 000 multiple choice questions in Registry format in the text and on the companion Evolve website provide a realistic exam experience Complete coverage helps you prepare for the ARRT Radiation Therapy Certification Exam Content review in outline and tabular format provides a concise recap of the material you need to know to succeed on the exam Exercises at the end of each section offer engaging active review opportunities

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Journal of Research of the National Bureau of Standards United States. National Bureau of Standards,1988

Electricity and Magnetism ,2005 **Atoms, Molecules & Elements: The Periodic Table Gr. 5-8** George

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