

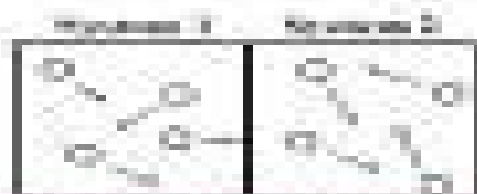
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Phase Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The particles in systems with more mass and change, instantly increased or not at temperature, increased. The Boltzmann-probability curves show that the average velocity of each of the distributions is increasing (less to large) and the average temperature is increasing.

3. Sketch each graph to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are moving faster than in System 2, so the temperature is higher. The particles in System 1 are moving faster than the particles in System 2, so the temperature is higher. The particles in System 1 are moving faster than the particles in System 2, so the temperature is higher.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are moving faster than the particles in System 2, so the temperature is higher. The particles in System 1 are moving faster than the particles in System 2, so the temperature is higher. The particles in System 1 are moving faster than the particles in System 2, so the temperature is higher.

Answer Key Unit 6 Thermochemistry

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