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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$ (B) Removing a small amount of $BaO(s)$ X (C) Adding He gas to the vessel X (D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

1996 Ap Chemistry Multiple Choice Answers

Neil D. Jespersen, Pamela Kerrigan



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[Children's Reading Comprehension and Assessment](#) Scott G. Paris,Steven A. Stahl,2005-03-23 The volume identifies critical issues and cutting edge research in assessing children s reading comprehension

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trend is expected to continue As a result a new item writing book is needed one that provides comprehensive coverage of both types of items and of the validity theory underlying them This book is an outgrowth of the author s previous book *Developing and Validating Multiple Choice Test Items* 3e Haladyna 2004 That book achieved distinction as the leading source of guidance on creating and validating selected response test items Like its predecessor the content of this new book is based on both an extensive review of the literature and on its author s long experience in the testing field It is very timely in this era of burgeoning testing programs especially when these items are delivered in a computer based environment Key features include Comprehensive and Flexible No other book so thoroughly covers the field of test item development and its various applications Focus on Validity Validity the most important consideration in testing is stressed throughout and is based on the Standards for Educational and Psychological Testing currently under revision by AERA APA and NCME Illustrative Examples The book presents various selected and constructed response formats and uses many examples to illustrate correct and incorrect ways of writing items Strategies for training item writers and developing large numbers of items using algorithms and other item generating methods are also presented Based on Theory and Research A comprehensive review and synthesis of existing research runs throughout the book and complements the expertise of its authors *Resources in Education* ,1997

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