

Solution

1. Answer (A, B, D)

n. factor of KMnO_4 in acidic medium = 5

In faintly basic/neutral medium = 3

2. Answer (A, D)

n.f. for $\text{C}_2\text{O}_4^{2-}$ or HC_2O_4^- with $\text{K}_2\text{Cr}_2\text{O}_7 = 2$

n.f. for HC_2O_4^- with $\text{HCl} = 1$ and $\text{C}_2\text{O}_4^{2-}$ with $\text{HCl} = 2$

3. Answer (C)



4. Answer (C)

Let the two metal carbonates be MCO_3 and $\text{M}'\text{CO}_3$. Let M be 13.6% by weight.

Weight of sample = 2.5 g

Weight of metal M = $\frac{13.6}{100} \times 2.5 = 0.34 \text{ g}$

Weight of CO_2 evolved = 1.33g

Moles of CO_2 evolved = $\frac{1.33}{44} = 0.0302$

$\therefore$ Weight of CO_3^{2-} in the sample = $0.0302 \times 60 = 1.812 \text{ g}$

$\therefore$ % of $\text{M}' = \frac{2.5 - 0.34 - 1.812}{2.5} \times 100 = 13.92\%$

5. Answer (A, B, C)



Eq. of H_2S = equivalent of I_2

$$\Rightarrow n_{\text{H}_2\text{S}} \times 2 = 2.40 \times 0.02$$

$$\Rightarrow n_{\text{H}_2\text{S}} = 2.40 \times 10^{-2} \text{ moles}$$

$\therefore$ % of S in steel sample

$$= \frac{2.4 \times 10^{-3} \times 32}{0.3 \times 100} \times 100$$

$$= 0.256\%$$

6. Answer (A, B, D)

For equimolar mixture of H_2SO_4 and SO_3

$$x + 80 + x \times 98 = 10$$

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