

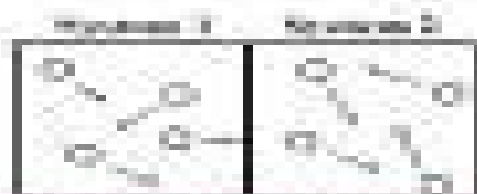
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The first curve of particles with maximum mass and average velocity contained an odd of temperatures, therefore, the Boltzmann-probability curve equal that the average velocity of each of the distributions is constant (that is why the maximum temperature is constant).

3. Sketch each graph to represent the relative distribution of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are more spread out, indicating a higher temperature. In System 2, the gas particles are more clustered together, indicating a lower temperature. The particles in System 1 are moving faster than the particles in System 2, which is why the particles in System 1 are more spread out than the particles in System 2.
- b. Are either distributed with kinetic energy for translation? Justify your answer in terms of molecular collisions.
All gases are distributed with kinetic energy for translation. In System 1, the gas particles are more spread out, indicating a higher temperature. In System 2, the gas particles are more clustered together, indicating a lower temperature. The particles in System 1 are moving faster than the particles in System 2, which is why the particles in System 1 are more spread out than the particles in System 2.

Answer Key Unit 6 Thermochemistry

Darrell D. Ebbing



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