

Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:



2) The rate law for the reaction



is $\text{rate} = k[\text{H}_2][\text{NO}]^2$. Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

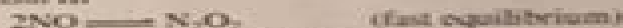
Mechanism I



Mechanism II



Mechanism III



3) Consider the reaction



From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:



Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of $\text{ClCO}_2\text{CCl}_3$ (mmHg)
0	15.76
181	12.64
513	8.73
1164	4.44

Answer Of Pre Lab Rates Of Reaction

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Answer Of Pre Lab Rates Of Reaction:

Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07
Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules. This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry. It not only highlights the key fundamental components of the coordination chemistry field, it also exemplifies the historical development of concepts in the field. In order to graduate as a chemistry major that fills the requirements of the American Chemical Society, a student needs to take a laboratory course in inorganic chemistry. Most professors who teach inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine, the environment, molecular biology, organic synthesis, and inorganic materials. **Chemistry**
Lemay, Prentice-Hall Staff, 1996 **Pre-lab Exercises for Experimental Organic Chemistry** Royston M. Roberts, John C. Gilbert, Stephen F. Martin, 1994 **Chemistry** McGraw-Hill Staff, 2001-07 **Pre-Lab Exercises for Modern Experimental Organic Chemistry** Royston M. Roberts, 1985 *Introductory Chemistry* Wendy Gloffke, Doris Kimbrough, 2002 This lab manual helps students develop data acquisition, organization, and analysis skills while teaching basic techniques. Students construct their own data tables, answer conceptual questions, and make predictions before performing experiments. They also have the opportunity to visualize and describe molecular level activity and explain the results. *Laboratory Experiments for General Chemistry* Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998
Chemical Investigations Nancy Konigsberg Kerner, 1986 **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach, presenting both micro, semi-micro, and macroscale techniques. The manual is easy to navigate with all relevant techniques found as they are needed. Cutting edge subjects such as HPLC, bioorganic chemistry, multistep synthesis, and more are presented in a clear and engaging fashion.
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laboratory professionals in the biotechnology field *Experimental Organic Chemistry* John C. Gilbert, Stephen F. Martin, 2006 This proven and well tested laboratory manual for organic chemistry students contains procedures for both miniscale also known as small scale and microscale users This lab manual gives students all the necessary background to enter the laboratory with the knowledge to perform the experiments with confidence For the microscale labs experiments were chosen to provide tangible quantities of material which can then be analyzed Chapters 1 2 introduce students to the equipment record keeping and safety of the laboratory Chapters 3 6 and 8 are designed to introduce students to laboratory techniques needed to perform all experiments In Chapters 7 and 9 through 20 students are required to use the techniques to synthesize compounds and analyze their properties In Chapter 21 students are introduced to multi step syntheses of organic compounds a practice well known in chemical industry In Chapter 23 students are asked to solve structures of unknown compounds The new chapter 24 introduces a meaningful experiment into the textbook that reflects the increasing emphasis on bioorganic chemistry in the sophomore level organic lecture course This experiment not only gives students the opportunity to accomplish a mechanistically interesting and synthetically important coupling of two amino acids to produce a dipeptide but also provides valuable experience regarding the role of protecting groups in effecting synthetic transformations with multiple functionalized molecules **Science Super Sleuths** Pam Walker, 1999 Develop students problem solving skills as they become detectives of science Fun filled experiments involve data tables graphs and conclusion drawing questions *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems *Science Spectrum* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 *Fundamentals of Chemistry in the Laboratory* Ralph A. Burns, 1999 **The Proceedings of the Iowa Academy of Science** Iowa Academy of Science, 1986 List of members in each volume **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** , 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **General Chemistry** Hunt, 1976-06-07 *Safety-Scale Lab Exp Biochem 2e* Spencer L. Seager, Michael R. Slabaugh, 1994-05 Chemical Engineering Education , 2002

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