

**Types of Chemical Reaction Worksheet**

Balance the reactions **1 to 6** and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

- |                                                                                                           |                                     |
|-----------------------------------------------------------------------------------------------------------|-------------------------------------|
| 1. $2 \text{NaBr} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaBr}_2 + 2 \text{NaOH}$                     | Reaction Type : single displacement |
| 2. $2 \text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow 3 (\text{NH}_4)_2\text{SO}_4$                       | Reaction Type : synthesis           |
| 3. $4 \text{C}_3\text{H}_8\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$    | Reaction Type : combustion          |
| 4. $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$          | Reaction Type : single displacement |
| 5. $\text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$ | Reaction Type : double displacement |
| 6. $\text{HBr} + 3 \text{Al}(\text{OH})_3 \rightarrow \text{H}_2\text{O} + 3 \text{AlBr}_3$               | Reaction Type : double displacement |

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

- |                                                                                                    |                                          |
|----------------------------------------------------------------------------------------------------|------------------------------------------|
| 7. $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$       | Reaction Type <b>double displacement</b> |
| 8. $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$              | Reaction Type <b>double displacement</b> |
| 9. $\text{C}_8\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$       | Reaction Type <b>combustion</b>          |
| 10. $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$                              | Reaction Type <b>single displacement</b> |
| 11. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$                                           | Reaction Type <b>decomposition</b>       |
| 12. $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$                                 | Reaction Type <b>synthesis</b>           |
| 13. $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + 2 \text{RbF}$           | Reaction Type <b>double displacement</b> |
| 14. $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$               | Reaction Type <b>single displacement</b> |
| 15. $\text{C}_3\text{H}_8\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ | Reaction Type <b>combustion</b>          |
| 16. $2 \text{C}_2\text{H}_5 + \text{Fe} \rightarrow \text{Fe}(\text{C}_2\text{H}_5)_2$             | Reaction Type <b>synthesis</b>           |
| 17. $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$                          | Reaction Type <b>single displacement</b> |
| 18. $2 \text{MgI}_2 + \text{Mn}(\text{SO}_3)_2 \rightarrow 2 \text{MgSO}_3 + \text{MnI}_4$         | Reaction Type <b>double displacement</b> |
| 19. $\text{O}_3 \rightarrow \text{O} + \text{O}_2$                                                 | Reaction Type <b>decomposition</b>       |
| 20. $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$                                          | Reaction Type <b>decomposition</b>       |

# Answers For Chemical Reactions

**Gill Murphy**



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