

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure elements and their substances. A metal is often used in metallurgy. It is used in chemical analysis or metallurgical analysis with chemical reactions.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to identify elements and compounds.
- Compare different metals and reactions.
- Describe the properties of different metals.

Module 14 Objectives

Lesson Question: How are metals used in chemical reactions?

Fill in the blanks.

1. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.
2. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.
3. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.

Chemical Reactions and Equations

For the first reaction, the reactants are iron and oxygen.

Chemical equation: $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

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Module 14

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