HOMOGENEOUS AQUEOUS SYSTEMS

Section Review

Objectives

- · Distinguish between a solvent and a solute
- · Describe what happens in the solution process
- Explain why all ionic compounds are electrolytes
- Demonstrate how to write the formula for a hydrate

Vocabulary

- aqueous solution
- solvent.
- solute

- solvation.
- electrolyte
- monelectrolyte

- strong electrolyte
- weak electrolyte
- hydrate.

Key Equation

Percent H₂O = mass of water x 100%

mass of hydrate

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water is a polar liquid and an excellent ______ for many substances. Aqueous solutions are _____ mixtures of ions or molecules in water. The solubility of a solute depends on solute-

solvent interactions. A good rule to remember is ____3__.

Substances that dissolve as ions are known as __4___ A
solute that is completely ionized in solution is a __5__
electrolyte. A weak electrolyte is only __6__ ionized. A solution
of an electrolyte will __7__ an electric current, whereas a

solution of a ____8 __ is nonconducting.

Many crystals are ______; they contain water of hydration.

In the process called ______ the water of hydration is lost from a

hydrate that is exposed to air.

3.

4.

5. _____

6.

7.

8. ____

10

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Answers For Homogeneous Aqueous Systems Section Review

American Chemical Society

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