

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1. (2)

2. (2)

3. (2)

4. (2)

5. (2)

6. (2) (2) (2) (2)

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2} \quad K_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CH}_4}) (P_{\text{CO}})^2}$$

gases: $\Delta n_{\text{gas}} = 2 + 1 - 3 = 0$ $Q_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CH}_4}) (P_{\text{CO}})^2} = \frac{(0.020)^2 (0.010)}{(0.010) (0.010)^2} = 4.0$ $K_p = 1.6$

Question 1

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.010 atm. The reaction is:



If the volume of an 8.00 L container at 1000°C where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: $[\text{H}_2] = 0.0025 \text{ mol/L}$
 Answer: $[\text{H}_2] = 0.0025 \text{ mol/L}$
 $[\text{H}_2] = 0.0025 \text{ mol/L}$
- Determine concentration of H_2S at equilibrium.
- Answer:
 What other species have chemical formula consisting of 2 hydrogen atoms and 1 sulfur?
- Answer:

	H_2	H_2S	$\text{S}_{(g)}$
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	0

1. Not balanced equation:

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