Chemical Formulas and Chemical Compounds

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

- In a Stock system name such as iron(III) sulfate, the Roman numeral tells us
 - (a) how many atoms of Fe are in one formula unit.
 - (b) how many sulfate ions can be attached to the iron atom.
 - (c) the charge on each Fe ion.
 - (d) the total positive charge of the formula unit.
- Changing a subscript in a correctly written chemical formula
 - (a) changes the number of moles represented by the formula.
 - (b) changes the charges on the other ions in the compound.
 - (c) changes the formula so that it no longer represents the compound it previously represented.
 - (d) has no effect on the formula.
- The explosive TNT has the molecular formula C₂H₂(NO₂)₃.
 - a. How many elements make up this compound?
 - b. How many oxygen atoms are present in one molecule of C₂H₃(NO₂)₃?
 - c. How many atoms in total are present in one molecule of C₇H₃(NO₂)₃?
 - d. How many atoms are present in a sample of 2.0 × 10²³ molecules of C₇H₅(NO₂)₃?
- 4. How many atoms are present in each of these formula units?
 - a. Ca(HCO₃)₂
 - b. C₁₂H₂₂O₁₁
 - c. Fe(ClO₂)₃
 - d. Fe(ClO₃)₂
- 5. ______ a. What is the formula for the compound dinitrogen pentoxide?
 - b. What is the Stock system name for the compound FeO?
 - c. What is the formula for sulfurous acid?
 - d. What is the name for the acid H₃PO₄?

Answers Review Chemical Formulas And Chemical Compounds

Stanley E. Manahan

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