

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\begin{array}{c} 21.5 \text{ g O}_2 \left| \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right| \frac{2 \text{ SO}_2}{3 \text{ O}_2} \left| \frac{3 \text{ S}}{\text{SO}_2} \right| \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S} \end{array}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\begin{array}{c} 4.6 \text{ mol Fe}(s) \left| \frac{3 \text{ CO}}{2 \text{ Fe}} \right| \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2 \end{array}$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\begin{array}{c} 153 \text{ g Ca}_3(\text{PO}_4)_2 \left| \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{214.02 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = \\ 76.8 \text{ g H}_2\text{SO}_4 \left| \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = \end{array}$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\begin{array}{c} 7.50 \text{ g H}_2\text{S} \left| \frac{1 \text{ mol H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \right| \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 10.89 \end{array}$$

$$\begin{array}{c} 12.75 \text{ g SO}_2 \left| \frac{1 \text{ mol SO}_2}{64 \text{ g SO}_2} \right| \frac{3 \text{ S}}{1 \text{ SO}_2} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 19.13 \end{array}$$

Ap Chemistry Multiple Choice Exam Answers

Albert A Gayle



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web 11 his head is buried near jacob the fates of esau and jacob merged even at death jacob passed on in egypt and his body was taken by his sons to be buried in the cave of machpelah in the land of canaan to their consternation however there was esau laying dubious claims to the last remaining plot in the cave

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web 33 jacob looked up and there was esau coming with his four hundred men so he divided the children among leah rachel and the two female servants 2 he put the female servants and their children in front leah and her children next and rachel and joseph in the rear 3 he himself went on ahead and bowed down to the ground seven times as he appr

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web the story focuses on esau s loss of his birthright to jacob and the conflict that ensued between their descendant nations because of jacob s deception of their aged and blind father isaac in order to receive esau s birthright blessing from isaac

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web the lord wanted jacob to have the birthright blessing because esau was not worthy of it jacob asked esau to trade him the birthright for some food esau agreed and traded his birthright to jacob genesis 25 23 29 34 hebrews 11 20 image

rebekah and isaac rebekah and isaac wanted what was best for their children

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