# Six Types of Chemical Reaction Worksheet

#### Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

- NaBr + 1 Ca(OH)<sub>2</sub> → 1 CaBr<sub>2</sub> + 2 NaOH
   Type of reaction: double displacement
- 2) 2 NH<sub>3</sub>+ 1 H<sub>2</sub>SO<sub>4</sub> → 1 (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> Type of reaction: synthesis
- 3) 4 C<sub>5</sub>H<sub>9</sub>O + 27 O<sub>2</sub> → 20 CO<sub>2</sub> + 18 H<sub>2</sub>O
  Type of reaction: combustion
- 4) 3 Pb + 2 H<sub>3</sub>PO<sub>4</sub> → 3 H<sub>2</sub> + 1 Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>
  Type of reaction: single displacement
- 5) 1 Li<sub>3</sub>N + 3 NH<sub>4</sub>NO<sub>3</sub> → 3 LiNO<sub>3</sub> + 1 (NH<sub>4</sub>)<sub>3</sub>N Type of reaction: double displacement
- 6) 3 HBr + 1 Al(OH)<sub>3</sub> → 3 H<sub>2</sub>O + 1 AlBr<sub>3</sub>

  Type of reaction: acid-base
- 7) What's the main difference between a double displacement reaction and an acid-base reaction? Acid-base reactions form water.
- 8) Combustion reactions always result in the formation of water. What other types of chemical reaction may result in the formation of water? Write examples of these reactions on the opposite side of this paper. Acid-base: HCl + NaOH → H₂O + NaCl Synthesis: 2 H₂ + O₂ → 2 H₂O Decomposition: NH₄OH → NH₃ + H₂O Single displacement: H₂ + 2 NaOH → 2 Na + 2 H₂O (not common)

# **Answere To Six Types Of Reaction Homework Worksheet**

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