

Experiment 2 - Gravimetric Analysis of a Chloride Salt (G)

Purpose:

Determining quantitative value of chloride in an unknown soluble salt by showing and explaining common methods in gravimetric analysis.

Theory:

- Reaction between silver ion and chloride ion to form silver chloride as an equation:



- AgCl solubility in water and the K_{sp} :

Silver chloride is very insoluble, so we can assume that the above reaction goes to completion.

$$K_{sp} = 1.6 \times 10^{-10}$$

- AgCl is precipitated with an acidic solution to reduce likelihood of acid ions interfering with results. These ions will form co-precipitates with silver in the acid, due to these co-precipitates, the results would be higher.
- To prevent the precipitate from going through the filtered paper, it is heated to coagulate it into a different texture that's easier to work with. Heating the precipitate makes it coagulate, therefore, it is harder for the precipitate to go through the filtered paper.
- Dry silver chloride decomposes into silver and chlorine in the presence of light:



- If photodecomposition takes place in air, results will be lower, since chlorine will be released in the form of gas. If photochemical decomposition occurs in the presence of excess silver ions in aqueous solution, it will result in higher results.

Analysis Of An Unknown Salt Lab Report

SB Merriam

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Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, 1994 *English in Analytical Chemistry. Communicating about Methods & Techniques. Книга для студента* Надежда Зорина, Александра Соболева, 2022-10-19

General Chemistry Hunt, 1976-06-07 **Scientific and Technical Aerospace Reports**, 1995-05

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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Energy Research Abstracts, 1992 **U.S. Government Research Reports**, 1964 **1998 Report on Attainment of the California Particulate Matter Standards in the Monterey Bay Region** Janet Brennan, 1999 **Chemistry** Karen Timberlake, 1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual. Each experiment encourages critical thinking with laboratory goals, discussion of related concepts, clear instructions, new pre-lab questions and comprehensive report pages. Forty-one experiments illustrate the basic principles of chemistry.

Analytical Chemistry for Technicians, Fourth Edition John Kenkel, 2013-08-13 Written as a training manual for chemistry-based laboratory technicians, this thoroughly updated fourth edition of the bestselling *Analytical Chemistry for Technicians* emphasizes the applied aspects rather than the theoretical ones. The book begins with classical quantitative analysis and follows with a practical approach to the complex world of sophisticated electronic instrumentation commonly used in real-world laboratories. Providing a foundation for the two key qualities: the analytical mindset and a basic understanding of the analytical instrumentation, this book helps prepare individuals for success on the job. Chapters cover sample preparation, gravimetric analysis, titrimetric analysis, instrumental analysis, spectrochemical methods such as atomic spectroscopy and UV-Vis and IR, molecular spectrometry, chromatographic techniques including gas chromatography and high performance liquid chromatography, electroanalytical methods and more. Incorporating an additional ten years of teaching experience since the publication of the third edition, the author has made significant updates and enhancements to the fourth edition. More than 150 new photographs and either new or reworked drawings spanning every chapter to assist the visual learner. A new chapter on mass spectrometry covering GC-MS, LC-MS, LC-MS/MS and ICP-MS. Thirteen new laboratory experiments. An introductory section before chapter 1 to give students a preview of general laboratory considerations, safety, laboratory notebooks and instrumental analysis. Additional end-of-chapter problems, expanded report-type questions and inclusion of relevant section headings in the Questions and Problems sections. Application Notes in each chapter. An appendix providing a glossary of quality assurance and good laboratory practice (GLP) terms.

Science Educator's Guide to Laboratory Assessment Rodney L. Doran, 2002 The book opens with an up-to-date discussion of assessment theory, research and uses. Then comes a wealth of sample assessment activities in biology, chemistry, physics and Earth science. Keyed to the National Science Education Standards, the activities include reproducible task sheets and scoring rubrics. All are ideal for helping students reflect on their own learning during science lab.

Report summaries United States. Environmental Protection Agency, 1983 *Nuclear*

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