

Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

- $$\overset{0}{\text{Mg}} + 2\overset{+1}{\text{H}}\overset{-1}{\text{Cl}} \rightarrow \overset{+2}{\text{Mg}}\overset{-2}{\text{Cl}}_2 + \overset{0}{\text{H}}_2$$

$$\overset{0}{\text{Mg}} \xrightarrow{-2e^-} \overset{+2}{\text{Mg}} \quad 2\overset{+1}{\text{H}} \xrightarrow{+2e^-} \overset{0}{\text{H}}_2$$
- $$\overset{0}{2\text{Fe}} + 3\overset{+3}{\text{V}}_2\overset{-2}{\text{O}}_3 \rightarrow \overset{+3}{\text{Fe}}_2\overset{-2}{\text{O}}_3 + 6\overset{+2}{\text{V}}\overset{-2}{\text{O}}$$

$$2\overset{0}{\text{Fe}} \xrightarrow{-6e^-} \overset{+3}{\text{Fe}}_2 \quad 3\overset{+3}{\text{V}}_2 \xrightarrow{+6e^-} 6\overset{+2}{\text{V}}$$
- $$\overset{+7}{\text{KMnO}}_4 + \overset{+5}{\text{KNO}}_2 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{SO}}_4 \rightarrow \overset{+2}{\text{Mn}}\overset{-2}{\text{SO}}_4 + 2\overset{+2}{\text{H}}_2\text{O} + \overset{+5}{\text{KNO}}_3 + \overset{-2}{\text{K}}_2\overset{-2}{\text{SO}}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnCl}_2 + \text{HCl} \rightarrow \text{CrCl}_3 + \text{SnCl}_4 + \text{KCl} + \text{H}_2\text{O}$$
- $$\text{KMnO}_4 + \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Cl}_2 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + \text{S} \rightarrow \text{SO}_2 + \text{KOH} + \text{Cr}_2\text{O}_3$$
- $$\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{K}_2\text{MnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$
- $$\text{Mn}(\text{NO}_3)_2 + \text{NaBiO}_3 + \text{HNO}_3 \rightarrow \text{HMnO}_4 + \text{Bi}(\text{NO}_3)_3 + \text{NaNO}_3 + \text{H}_2\text{O}$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$

Analyzing Oxidation Reduction Reactions Answer Key

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A text book of PHARMACEUTICAL ANALYSIS-I Dr. Alok kumar Dash, 2025-06-03 A Pharmaceutical Analysis 1 Theory Practical textbook for B Pharm first semester focuses on the fundamental principles and techniques used to analyze pharmaceutical drugs ensuring their purity quality and safety The book typically covers topics like titration methods qualitative and quantitative analysis and the basics of analytical instrumentation as mandated by the Pharmacy Council of India PCI regulations It aims to provide students with a strong foundation in analytical chemistry relevant to the pharmaceutical industry

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Publications of the National Bureau of Standards ... Catalog United States. National Bureau of Standards, 1978

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