Science

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(Chapter - 4) (Structure of the Atom) (Exercises)
(Class - IX)

Question 1:

Compare the properties of electrons, protons and neutrons.

Answer 1:

Electron	Proton	Neutron
(i) Electron are present outside the nucleus of an atom.	Proton are present in the nucleus of an atom.	(i) Neutron are present in the nucleus of an atom.
(ii) Electron are negatively charged	(ii) Proton are positively charged.	(ii) Neutron are neutral.
(iii) The mass of an electron is considered to negligible.	(iii) The mass of a proton is approximately 2000 times as the mass of an electron	

Question 2:

What are the limitations of J.J. Thomson's model of the atom?

EAnswer 2:

According to J.J. Thomson's model of an atom, an atom consists of a positively charged sphere with electrons embedded in it. However, it was later found that the positively charged particles reside at the center of the atom called the nucleus, and the electrons revolve around the nucleus.

Question 3:

What are the limitations of Rutherford's model of the atom?

Answer 3:

According to Rutherford's model of an atom, electrons revolve around the nucleus in fixed orbits. But, an electron revolving in circular orbits will not be stable because during revolution, it will experience acceleration. Due to acceleration, the electrons will lose energy in the form of radiation and fall into the nucleus. In such a case, the atom would be highly unstable and collapse.

Ouestion 4:

Describe Bohr's model of the atom.

Answer 4:

Bohr's model of the atom Niels Bohr proposed the following postulates regarding the model of the atom.

- > Only certain orbits known as discrete orbits of electrons are allowed inside the atom.
- While revolving in these discrete orbits, the electrons do not radiate energy. These discrete orbits or shells are shown in the following diagram.



The first orbit (i.e., for n=1) is represented by letter K. Similarly, for n=2, it is L- shell, for n=3, it is M- shell and for n=4, it is N- shell. These orbits or shells are also called energy levels.

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Table of Contents Answer Of Ncert Class 9th

- 1. Understanding the eBook Answer Of Ncert Class 9th
 - o The Rise of Digital Reading Answer Of Ncert Class 9th
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer Of Ncert Class 9th
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Of Ncert Class 9th
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer Of Ncert Class 9th
 - Personalized Recommendations
 - Answer Of Ncert Class 9th User Reviews and Ratings
 - Answer Of Ncert Class 9th and Bestseller Lists
- 5. Accessing Answer Of Ncert Class 9th Free and Paid eBooks
 - Answer Of Ncert Class 9th Public Domain eBooks
 - Answer Of Ncert Class 9th eBook Subscription Services
 - Answer Of Ncert Class 9th Budget-Friendly Options
- 6. Navigating Answer Of Ncert Class 9th eBook Formats

- o ePub, PDF, MOBI, and More
- Answer Of Ncert Class 9th Compatibility with Devices
- Answer Of Ncert Class 9th Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Of Ncert Class 9th
 - Highlighting and Note-Taking Answer Of Ncert Class 9th
 - Interactive Elements Answer Of Ncert Class 9th
- 8. Staying Engaged with Answer Of Ncert Class 9th
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Of Ncert Class 9th
- 9. Balancing eBooks and Physical Books Answer Of Ncert Class 9th
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Of Ncert Class 9th
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Of Ncert Class 9th
 - Setting Reading Goals Answer Of Ncert Class 9th
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Of Ncert Class 9th
 - Fact-Checking eBook Content of Answer Of Ncert Class 9th
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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