

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure elements and their compounds. A metal is often used in metallurgy. It is used in manufacturing or metallurgical analysis with chemical techniques.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical equations for chemical reactions and mixtures.
- Calculate the mass of a substance.
- Calculate the mass of a substance.

Module 14 Objectives

Learn to use the law of conservation of mass to balance chemical reactions.

Fill in the blanks:

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ (Catalyst)
2. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
3. $\text{Ca} + \text{Mg} \rightarrow \text{Ca} + \text{Mg}$ (Catalyst)

Chemical Reactions and Equations

For the first reaction, the law of conservation of mass:

Conservation of mass - the total mass of the reactants is equal to the total mass of the products.

Reaction equation: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Reaction equation: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Reaction equation: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Module 14

Learn to use the law of conservation of mass to balance chemical reactions.

Fill in the blanks:

1. 1 mole of H_2 is equal to $\frac{1}{2}$ mole of H_2 .
2. The mass of the reactants is equal to the mass of the products. (Conservation of mass)

Answers For Chemistry Module 14

Siddappa N.Byrareddy



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