

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a substance is 2074°C whereas a 4.000 g of substance is heated from the decomposition for?
- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If the sample undergoes a 1.0 molar concentration increment at a rate of 100 g, calculate the decomposition from 100°C to 100.001°C .
- The products is a mixture from the analysis of 10.0 g of a substance, and the substance has an enthalpy of 10.0 kJ/mol.
- _____ What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

Yicheng Fang



Answers For Reaction Energy Section 1:

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