

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₃.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₃ produced with 95g of Fe (SO₄)₃

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

Answers For Study Guide Mass And Moles

John McMurry, Susan McMurry



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
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Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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Table of Contents Answers For Study Guide Mass And Moles

1. Understanding the eBook Answers For Study Guide Mass And Moles
 - The Rise of Digital Reading Answers For Study Guide Mass And Moles
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Study Guide Mass And Moles
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Study Guide Mass And Moles
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Study Guide Mass And Moles
 - Personalized Recommendations
 - Answers For Study Guide Mass And Moles User Reviews and Ratings
 - Answers For Study Guide Mass And Moles and Bestseller Lists
5. Accessing Answers For Study Guide Mass And Moles Free and Paid eBooks
 - Answers For Study Guide Mass And Moles Public Domain eBooks
 - Answers For Study Guide Mass And Moles eBook Subscription Services
 - Answers For Study Guide Mass And Moles Budget-Friendly Options
6. Navigating Answers For Study Guide Mass And Moles eBook Formats

- ePub, PDF, MOBI, and More
- Answers For Study Guide Mass And Moles Compatibility with Devices
- Answers For Study Guide Mass And Moles Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Study Guide Mass And Moles
 - Highlighting and Note-Taking Answers For Study Guide Mass And Moles
 - Interactive Elements Answers For Study Guide Mass And Moles
- 8. Staying Engaged with Answers For Study Guide Mass And Moles
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Study Guide Mass And Moles
- 9. Balancing eBooks and Physical Books Answers For Study Guide Mass And Moles
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Study Guide Mass And Moles
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Study Guide Mass And Moles
 - Setting Reading Goals Answers For Study Guide Mass And Moles
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Study Guide Mass And Moles
 - Fact-Checking eBook Content of Answers For Study Guide Mass And Moles
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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