

Section A

Marks

1. (a) (i) ^{14}C is a β -particle emitter and the product of decay, ^{14}N , is not radioactive

1



(1)

When a living organism dies, intake of C stops while decay of ^{14}C continues

1

Therefore, $^{14}\text{C} : ^{12}\text{C}$ ratio in dead living organisms decreases. The age of an ancient biological specimen can be estimated by comparing its activity to that of living organisms alive.

(ii) $N = N_0 e^{-\lambda t}$

$$\ln\left(\frac{1}{2}\right) = -\lambda (5730)$$

1

$$\lambda = 1.21 \times 10^{-4} \text{ y}^{-1}$$

1

$$\ln\left(\frac{9.6}{15.3}\right) = -1.21 \times 10^{-4} t$$

$$t = 3850 \text{ years}$$

1

 (5)

- (b) (i) The sample is vaporized in the mass spectrometer and is bombarded by fast moving electrons to give positive ions

1



(1)

The ions are accelerated by an electric field.

1

The accelerated ions are then deflected by a magnetic (electric) field.

1

The mass spectrometer is operated at low pressure ($\sim 10^{-7}$ atm) to prevent ions from colliding with other particles

1

The lighter the positive ion, the greater is the deflection

1

By varying the strength of the deflecting magnetic field, ions of a particular mass / charge (m/e) ratio are brought to the ion detector.

1

Determination of the relative atomic mass of a noble gas:

The current intensity produced by cations formed from an isotope is proportional to the relative abundance of the isotope.

1

relative atomic mass of the noble gas

$$= \Sigma (\text{isotopic mass} \times \text{relative abundance})$$

 (8)

Advanced Higher Chemistry Paper 2001 Marking Scheme

Susan Sharpless Smith



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