Balancing redox equations

Example

Balance the equation showing the oxidation of Fe²⁺ ions to Fe³⁺ ions by dichromate ions (Cr₂O₇²⁻) in an acidic medium

$$Fe^{2+} + Cr_2O_7^{2-} \rightarrow Fe^{3+} + Cr^{3+}$$

Step 1: Identify oxidising and reducing agents and write half reactions

Fe²⁺: +2
$$Cr_2O_7^{2-}$$
: $2Cr + 7(-2) = -2$
Fe³⁺: +3 $2Cr = 12$
 Cr^{3+} : +3 $Cr = +6$

$$Fe^{2+} \rightarrow Fe^{3+}$$

Iron loses one electron, i.e. it is oxidised. It donates one electron to dichromate, i.e. it is the reducing agent.

$$Cr_2^{6+} \rightarrow Cr^{3+}$$

Chromium gains three electrons, i.e. it is reduced. It gains its electrons from iron, i.e. it is the oxidising agent.

Half equation: Reduction reaction: Cr26+ + 3e- → Cr3+

Analyzing Oxidation Reduction Reactions Answers

Henry Paul Talbot

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Table of Contents Analyzing Oxidation Reduction Reactions Answers

- 1. Understanding the eBook Analyzing Oxidation Reduction Reactions Answers
 - The Rise of Digital Reading Analyzing Oxidation Reduction Reactions Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Analyzing Oxidation Reduction Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analyzing Oxidation Reduction Reactions Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Analyzing Oxidation Reduction Reactions Answers
 - Personalized Recommendations
 - Analyzing Oxidation Reduction Reactions Answers User Reviews and Ratings
 - Analyzing Oxidation Reduction Reactions Answers and Bestseller Lists
- 5. Accessing Analyzing Oxidation Reduction Reactions Answers Free and Paid eBooks

- Analyzing Oxidation Reduction Reactions Answers Public Domain eBooks
- Analyzing Oxidation Reduction Reactions Answers eBook Subscription Services
- Analyzing Oxidation Reduction Reactions Answers Budget-Friendly Options
- 6. Navigating Analyzing Oxidation Reduction Reactions Answers eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Analyzing Oxidation Reduction Reactions Answers Compatibility with Devices
 - Analyzing Oxidation Reduction Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analyzing Oxidation Reduction Reactions Answers
 - Highlighting and Note-Taking Analyzing Oxidation Reduction Reactions Answers
 - Interactive Elements Analyzing Oxidation Reduction Reactions Answers
- 8. Staying Engaged with Analyzing Oxidation Reduction Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analyzing Oxidation Reduction Reactions Answers
- 9. Balancing eBooks and Physical Books Analyzing Oxidation Reduction Reactions Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Analyzing Oxidation Reduction Reactions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analyzing Oxidation Reduction Reactions Answers
 - Setting Reading Goals Analyzing Oxidation Reduction Reactions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analyzing Oxidation Reduction Reactions Answers
 - Fact-Checking eBook Content of Analyzing Oxidation Reduction Reactions Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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