

Answers Chemistry Review Packet

1. Organisms are made of the same chemical elements as nonliving things. One important difference is the complexity of organization in living things when compared to nonliving. The atoms of living things are organized into small molecules like sugars, amino acids, fatty acids, and nucleic acids. These small molecules are linked together to form macromolecules, and the macromolecules are organized into larger structures called macromolecular assemblies. Examples of macromolecular assemblies include ribosomes, membranes and chromatin.
2. The six most abundant elements (in decreasing order of abundance) are Oxygen, Carbon, Hydrogen, Nitrogen, Calcium, Phosphorus.

Mnemonic →

Trace elements are those required by organisms in only minute quantities (less than 0.01% of human body weight). These are B, Cr, Co, Cu, F, I, Fe, Mn, Mo, Se, Si, Sn, V, Zn. (Note: It is not necessary to memorize the names or symbols of these trace elements.)

Many trace elements are enzyme cofactors (e.g. Cu, Zn, Mn, Mg, Mo). Other functions include serving as electrolytes (nerve function), in osmotic balance, and in acid-base balance. Check out Table 41.2 Mineral requirements of Humans on page 852 of Campbell (I love that book!). Don't memorize the table, just look at it. It's cool!

3. Atoms have a nucleus at the center which is surrounded by an electron cloud. The electron cloud is organized into a series of seven progressively larger concentric shells (somewhat like an onion).

The ratio of the diameter of the electron cloud to the diameter of the nucleus is

$$\frac{\text{Electron cloud}}{\text{Nucleus}} = \frac{10^{-8} \text{ cm}}{10^{-13} \text{ cm}} = 10^5$$

Thus the diameter of the electron cloud is 100,000 times larger than the diameter of the nucleus.

If the nucleus were as large as a baseball (with a radius of about 3.5 cm), the radius of the electron cloud would be ~2 miles.

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Jiyuan Zhang



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Computational Chemistry: Reviews Of Current Trends, Vol. 5 Ermanno Gianinetti, Jozek S Kwiatkowski, Jerzy Leszczynski, Piotr Piecuch, Jiri Sponer, Thanh N Truong, Charles A Weatherford, 2000-10-09 This volume comprises six chapters which explore the development and applications of the methods of computational chemistry The first chapter is on new developments in coupled cluster CC theory The homotopy method is used to obtain complete sets of solutions of nonlinear CC equations The correspondence between multiple solutions to the CCSD CCSDT and full CI equations is established and the applications of the new approach in modeling molecular systems are discussed The second chapter reviews the computational theory for the time dependent calculations of a solution to the Schrödinger equation for two electrons and focuses on the development of propagators to the solution The next chapter features a discussion on a new self consistent field for molecular interactions SCF MI scheme for modifying Roothaan equations in order to avoid basis set superposition errors BSSE This method is especially suitable for computations of intermolecular interactions Details of the theory along with examples of applications to nucleic acid base pair complexes are given This chapter is well complemented by the following chapter which reports the current status of computational studies of aromatic stacking and hydrogen bonding interactions among nucleic acid bases The next chapter reveals the possibility of calculating the kinetics of chemical reactions in biological systems from the first principles The last chapter reviews the results of rigorous ab initio studies of the series of derivatives of methane silane and germane The presented molecular and vibrational parameters complement experimental data for these systems In addition the theoretical approach allows the prediction of the effects of halogeno substitutions on their structures and properties

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