

Chapter 1 - Chemical Foundations

1. Consider the element indium, atomic number 49, atomic mass 114.8 g. The nucleus of an atom of indium-112 contains:

- a. 49 protons, 63 neutrons, 49 electrons.
- b. 49 protons, 49 neutrons.
- c. 49 protons, 49 alpha particles.
- d. 49 protons, 63 neutrons.
- e. 49 protons, 112 neutrons.

ANSWER:

d

POINTS:

1

2. A hypothetical element consists of two isotopes of masses 78.95 amu and 80.95 amu with abundances of 32.9% and 67.1%, respectively. What is the average atomic mass of this element?

- a. 79.95 amu
- b. 78.95 amu
- c. 79.6 amu
- d. 80.3 amu
- e. 80.95 amu

ANSWER:

d

POINTS:

1

3. Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29 (28.976 amu, 4.67% abundance), and X-30 (29.974 amu, 3.10% abundance). Calculate the atomic weight of X.

- a. 28.09 amu
- b. 48.63 amu
- c. 27.16 amu
- d. 28.97 amu
- e. 86.93 amu

ANSWER:

a

POINTS:

1

4. The atomic mass of rhenium is 186.2. Given that 37.1% of natural rhenium is rhenium-185, what is the other stable isotope?

- a. $^{183}_{75}\text{Re}$
- b. $^{187}_{75}\text{Re}$
- c. $^{186}_{75}\text{Re}$
- d. $^{181}_{75}\text{Re}$
- e. $^{190}_{75}\text{Re}$

ANSWER:

b

POINTS:

1

5. Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29

Answers For Chapter One Chemical Foundations Exercises

Santosh Kumar Srivastava



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Spins in Chemistry Roy McWeeny,2013-01-18 Profile of evolution of the spin concept from its role in quantum mechanics to its assimilation into the field of chemistry examines spin and valence spin Hamiltonians more 1970 edition

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understand as opposed to explanations of complex chemical theory Mark's Basic emphasizes clinical correlations throughout the text and links biochemical concepts to physiology and pathophysiology using patient vignettes as the context These specific and memorable mock patient cases are followed throughout the chapter to pose questions illustrate core concepts and help students remember and apply biochemical principles within the context of clinical practice **Congressional Record** United States. Congress, 1919 A Short History of Chemistry James Riddick Partington, 1989-01-01 This classic exposition explores the origins of chemistry alchemy early medical chemistry nature of atmosphere theory of valency laws and structure of atomic theory and much more **The Development of Modern Chemistry** Aaron J. Ihde, 1984-01-01 From ancient Greek theory to the explosive discoveries of the 20th century this authoritative history shows how major chemists their discoveries and political economic and social developments transformed chemistry into a modern science 209 illustrations 14 tables Bibliographies Indices Appendices

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