

PHYSICAL KINETIC MOLECULAR THEORY

Learning Objectives

- to identify theoretical differences in how gases are represented in the kinetic theory and in real-world situations;
- to explain the relationship of the postulates of the kinetic molecular theory;
- to explain how molecular theory affects macroscopic behavior;
- to represent quantitative information using graphical or tabular representation.

Module 1: Representation of Gases in Different Phases

The following figure illustrates the representation of a gas in three different phases: solid, liquid, and gas.

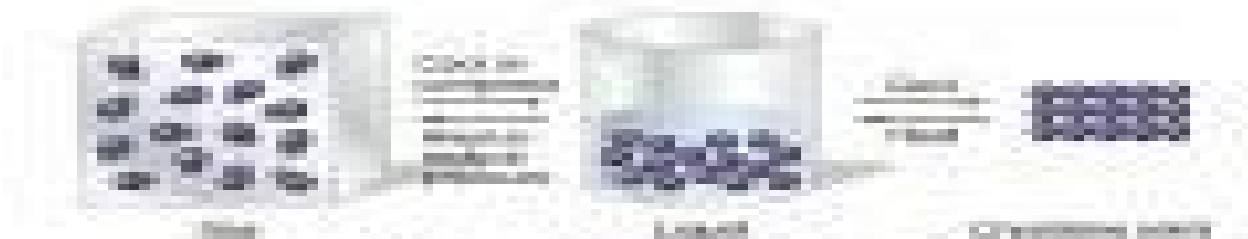


Figure 1: Representation of a gas in three different phases: solid, liquid, and gas.

Real-world gases are represented by different equations of state and properties. The kinetic molecular theory is a simplified model of gas behavior.

Key Concepts

- The kinetic molecular theory of gases is based on the following assumptions:
 - Gases are composed of particles that are in constant motion.
 - The particles are small compared to the distance between them.
 - The particles are in constant motion and collide with each other and the walls of the container.

Table 1: Kinetic Molecular Theory of Gases

	SOLID	LIQUID	GAS
Particle arrangement	Particles are packed closely together in a regular, repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.	Particles are packed closely together in a disordered, non-repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.	Particles are spread out in a disordered, non-repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.
Particle arrangement in the container	Particles are packed closely together in a regular, repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.	Particles are packed closely together in a disordered, non-repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.	Particles are spread out in a disordered, non-repeating pattern. The particles are in constant motion and collide with each other and the walls of the container.

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