

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{54}_{26}\text{Fe}$	B2																
3 (b)	$^{24}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th><th>charge</th></tr></thead><tbody><tr><td>^{17}O</td><td>8</td><td>9</td><td>10</td><td>2-</td></tr><tr><td>^{27}Al</td><td>13</td><td>14</td><td>10</td><td>3+</td></tr></tbody></table>		protons	neutrons	electrons	charge	^{17}O	8	9	10	2-	^{27}Al	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
^{17}O	8	9	10	2-														
^{27}Al	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th></tr></thead><tbody><tr><td>^{32}S</td><td>16</td><td>16</td><td>16</td></tr><tr><td>$^{32}\text{S}^{2-}$</td><td>16</td><td>16</td><td>18</td></tr></tbody></table>		protons	neutrons	electrons	^{32}S	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
^{32}S	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2(\text{SO}_4)_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

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**Douglas Buchanan, James Melrose, A. R.
Ross**

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 existing issue resolution processes and compliance with whs laws and related documentation

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 appropriate timeframes 3 2 contribute to workplace meetings workplace inspections or other consultative activities 3 3
 provide assistance to workgroup members or other stakeholders to contribute to workplace safety

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web 2 contribute to whs consultation 2 1 contribute to facilitating whs consultation and participation processes according to whs laws 2 2 identify and report barriers to effective whs consultation and participation processes to required personnel according to organisational policies and procedures 2 3

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web 1 1 contribute to determining whs regulatory requirements and standards that apply to monitoring 1 2 consult with relevant individuals and or parties to identify agents and or conditions to be measured according to organisational policies and procedures cooperates with others as part of whs monitoring processes and contributes to

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web 1 contribute to the development of whs issue resolution processes 1 1 identify relevant standards whs laws and organisational policies procedures processes and systems that apply to the resolution of whs issues 1 2 review sufficiency of existing issue resolution processes and compliance with whs laws and related documentation

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web contribute to whs risk assessment 4 1 identify individuals and or parties at risk of exposure to hazards and determine the nature severity and likelihood of potential harm 4 2 contribute to applying tools techniques and processes to identified

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