## Analytical Chemistry MCQs: TYBSC SEM III

 The solubility product K<sub>sp</sub> of sparingly soluble Ag<sub>2</sub>CrO<sub>4</sub> is 4\*10<sup>-12</sup>. The molar solubility of the salt is

- (a) 1.0\*10\* moles/lit
- (b) 2.0\*104 moles/lit
- (c) 1.0\*10-5 moles/lit
- (d) 1.2\*10-12 moles/lit

 Solubility product of AgCl is 1.44\*10<sup>-4</sup> moles/lit at 373K. The solubility of the AgCl is boiling of water will be....

- (a) 0.72\*104 moles/lit
- (b) 0.72\*10-2 moles/lit
- (c) 1.20\*10-2 moles/lit
- (d)1.20\*10<sup>-4</sup> moles/lit
- 10. Precipitation takes place when the ionic product...
- (a) Is zero
- (b) Exceeds the solubilty product
- (c) Equal to solubilty product
- (d) Is less than the solubility product
- 11. The electomotive force is expressed In the unit of .....
- (a) Ohm
- (b) Volt
- (c) Ampere
- (d) Colulomb
- 12. If the cell can provide electrical energy to an external system, it is called as
- (a) Galvanic cell
- (b) Electrolytic cell
- (c) Concentration cell
- (d) Chemical cell
- The current per unit area of electrode surface is known as....
- (a) Current efficiency
- (b) Current yield
- (e) Current density
- (d) None of these

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