

## CHAPTER 10: Reactions

**Reaction Energy****SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of  $\Delta H_f^\circ$  is \_\_\_\_\_.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- \_\_\_\_\_  $\Delta H_{\text{f}}^\circ$  of a substance is  $2074^\circ\text{C}$ . Absorption of 4.000 J of energy will result in the decomposition of \_\_\_\_\_.
- \_\_\_\_\_ A substance has a specific heat of  $0.001 \text{ J/g}^\circ\text{C}$ . If a sample weighing  $10.1 \text{ g}$  is heated from  $0^\circ\text{C}$  to  $100^\circ\text{C}$ , it will absorb \_\_\_\_\_ J of energy.
- The product is a molecular form of oxygen, and the reactant has an enthalpy of  $0 \text{ kJ/mol}$ .  
\_\_\_\_\_ Is this a  $\Delta H_f^\circ$  or  $\Delta H_d^\circ$  reaction?

# Answers For Reaction Energy Section 1

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