

IONIC & COVALENT BONDS

Name:

Date:

Period:

An _____ is the attraction between two oppositely charged ions. The chemical bond formed when two atoms share electrons is called a _____. Covalent bonds usually form between atoms of _____ and ionic bonds usually form when a _____ combines with a _____.

Activity A: How Do Ions Form?

Procedure:

1. Place three pairs of checkers (three red and three black) on your desk. The red represent electrons and the black represent protons.
2. Place nine pairs of checkers (nine red and nine black) in a separate group on your desk.
3. Move a red checker from the smaller group to the larger group.
4. Count the number of positive charges (protons) and negative charges (electrons) in each group.
5. Now sort the checkers into a group of four pairs and a group of eight pairs. Repeat Steps 3 and 4, this time moving two red checkers from the smaller group to the larger group.

Think it over:

1. What was the total charge on each group before you moved the red checkers (electrons)?
2. What was the charge on each group after you moved the checkers?
3. Based on this activity what do you think happens to the charge on an atom when it loses electrons?
4. Based on this activity what do you think happens to the charge on an atom when it gains electrons?

Activity B: Interpreting Data About Ions

Procedure:

1. Use the periodic table and the chart to the right to identify the charges of the ions in each ionic compound listed below.
2. Write the formulas for each compound.
 - a. sodium bromide;
 - b. lithium oxide;
 - c. magnesium sulfide;
 - d. aluminum fluoride;
 - e. potassium nitrate;
 - f. ammonium chloride

Ions and Their Charges		
Name	Charge	Symbol or Formula
Lithium	1+	Li^+
Sodium	1+	Na^+
Potassium	1+	K^+
Ammonium	1+	NH_4^+
Calcium	2+	Ca^{2+}
Magnesium	2+	Mg^{2+}
Aluminum	3+	Al^{3+}
Fluoride	1-	F^-
Chloride	1-	Cl^-
Iodide	1-	I^-
Bicarbonate	1-	HCO_3^-
Nitrate	1-	NO_3^-
Oxide	2-	O^{2-}
Sulfide	2-	S^{2-}
Carbonate	2-	CO_3^{2-}
Sulfate	2-	SO_4^{2-}
Phosphate	3-	PO_4^{3-}

Think it over:

1. How do you know how many of each ion to write in the formula?

Answers Covalent Bonding Vocabulary

Zhe Xu



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Table of Contents Answers Covalent Bonding Vocabulary

1. Understanding the eBook Answers Covalent Bonding Vocabulary
 - The Rise of Digital Reading Answers Covalent Bonding Vocabulary
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Covalent Bonding Vocabulary
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Covalent Bonding Vocabulary
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Covalent Bonding Vocabulary
 - Personalized Recommendations
 - Answers Covalent Bonding Vocabulary User Reviews and Ratings
 - Answers Covalent Bonding Vocabulary and Bestseller Lists
5. Accessing Answers Covalent Bonding Vocabulary Free and Paid eBooks
 - Answers Covalent Bonding Vocabulary Public Domain eBooks
 - Answers Covalent Bonding Vocabulary eBook Subscription Services
 - Answers Covalent Bonding Vocabulary Budget-Friendly Options
6. Navigating Answers Covalent Bonding Vocabulary eBook Formats

- ePub, PDF, MOBI, and More
- Answers Covalent Bonding Vocabulary Compatibility with Devices
- Answers Covalent Bonding Vocabulary Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Covalent Bonding Vocabulary
 - Highlighting and Note-Taking Answers Covalent Bonding Vocabulary
 - Interactive Elements Answers Covalent Bonding Vocabulary
- 8. Staying Engaged with Answers Covalent Bonding Vocabulary
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Covalent Bonding Vocabulary
- 9. Balancing eBooks and Physical Books Answers Covalent Bonding Vocabulary
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Covalent Bonding Vocabulary
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Covalent Bonding Vocabulary
 - Setting Reading Goals Answers Covalent Bonding Vocabulary
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Covalent Bonding Vocabulary
 - Fact-Checking eBook Content of Answers Covalent Bonding Vocabulary
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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