

1 Complete the word equations for the reactions of the Group 7 elements with metals.

- a iron + chlorine → iron chloride
 b iron + bromine → iron bromide
 c sodium + iodine → sodium iodide
 d potassium + bromine → potassium bromide
 e lithium + chlorine → lithium chloride
 f sodium + chlorine → sodium chloride

2 Make up four sentences using the phrases in the table. Each sentence must include a phrase from each column.

There are	weak forces	between the two atoms of a chlorine molecule.
	strong forces	between a chlorine molecule and its neighbours.
		between the two atoms of an iodine molecule.
		between an iodine molecule and its neighbours.

*There are weak forces between the two atoms of a chlorine molecule.
 There are strong forces between a chlorine molecule and its neighbours.
 There are weak forces between the two atoms of an iodine molecule.
 There are strong forces between an iodine molecule and its neighbours.*

3 For the sentences A–H below:

- Write **1** next to the sentences that are true for Group 1.
 - Write **7** next to the sentences that are true for Group 7.
 - Write **B** next to the sentences that are true for both Groups 1 and 7.
- A Going down this group, boiling point increases.
 B Going up this group, the reactions of the elements get more vigorous.
 C Going down this group, the number of protons in atoms of the elements increases.
 D Atoms of the elements in this group form molecules made up of two atoms.
 E Going up this group, the colours of the elements get lighter.
 F The elements at the top of this group are gases.
 G The elements in this group conduct electricity.
 H All the atoms of the elements in this group have the same number of electrons in the outermost shell.

7
1
B
1
5
7
2, 7
B

12 Write **M** next to the pair of elements that you predict will react most vigorously.
 Write **L** next to the pair of elements that you predict will react least vigorously.

- a sodium and iodine
 b lithium and iodine
 c potassium and fluorine
 d potassium and bromine
 e sodium and bromine

L
L
M
M
L

Answers For A Ls Chemistry

Padhraic Smyth



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